

Name _____ Date _____ Class _____

CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

States of Matter

Section 13.1 Gases

In your textbook, read about the kinetic molecular theory.

Complete each statement.

1. The kinetic molecular theory describes the behavior of gases in terms of particles in motion.
2. The kinetic molecular theory makes the following assumptions:
 - a. In a sample of a gas, the volume of the gas particles themselves is very small compared to the volume of the sample.
 - b. Because gas particles are far apart, there are no significant attractive or repulsive forces between gas particles.
 - c. Gas particles are in constant and random motion.
 - d. The collisions between gas particles are elastic; that is, no kinetic energy is lost.
3. The kinetic energy of a particle is represented by the equation $KE = \frac{1}{2}mv^2$.
4. Temperature is a measure of the average kinetic energy of the particles in a sample of matter.

In your textbook, read about explaining the behavior of gases.

For each statement below, write true or false.

- | | |
|--------------|--|
| <u>True</u> | 1. Gases are less dense than solids because there is a lot of space between the particles of a gas. |
| <u>True</u> | 2. The random motion of gas particles causes a gas to expand until it fills its container. |
| <u>False</u> | 3. The density of a gas decreases as it is compressed. |
| <u>True</u> | 4. A gas can flow into a space occupied by another gas. |
| <u>True</u> | 5. The diffusion of a gas is caused by the random motion of the particles of the gas. |
| <u>False</u> | 6. Lighter gas particles diffuse less rapidly than do heavier gas particles. |
| <u>True</u> | 7. During effusion, a gas escapes through a tiny opening into a vacuum. |
| <u>False</u> | 8. Graham's law of effusion states that the rate of effusion for a gas is directly related to the square root of its molar mass. |

Name _____ Date _____ Class _____

CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

Section 13.2 continued

In your textbook, read about gas pressure.

Circle the letter of the choice that best completes the statement or answers the question.

13. Pressure is defined as force per unit c. area.
14. What is an instrument designed to measure atmospheric pressure? c. sphygmomanometer
15. The height of the liquid in a barometer is affected by all of the following EXCEPT the a. altitude.
16. The pressure of the gas in a manometer is directly related to which of the following quantities? b. height of the mercury column in the closed end arm
17. One atmosphere is equal to a pressure of b. 101.3 kPa.
18. The partial pressure of a gas depends on all of the following EXCEPT the b. size of the container.
19. The pressure of a sample of air in a manometer is 102.1 kPa. What is the partial pressure of nitrogen (N_2) in the sample if the combined partial pressures of the other gases is 22.4 kPa? b. 79.7 kPa

Use the figure to answer the following questions.

20. What instrument is illustrated in the figure? barometer
21. Who invented this instrument? Evangelista Torricelli
22. What are the two opposing forces that control the height of the mercury in the column? gravity and atmospheric pressure
23. What does it mean when the level of mercury drops in the column? The atmospheric pressure has increased.



Chemistry Matter Change Study Guide Answers 1

Arshad Iqbal



Chemistry Matter Change Study Guide Answers 1:

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement Study Guide to Accompany Calculus for the Management, Life, and Social Sciences Clyde Metz, 1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences Chemistry Equations & Answers (Speedy Study Guide) Speedy Publishing, 2014-11-26 Chemistry can be one of the more difficult subjects in school mainly because memorization is crucial It can be hard to memorize tons of facts and equations A chemistry equations and answers guide does help students some more than others You can refer to the study guide to determine which formula is suitable to use for the question at hand The more savvy student might consider one or more ways to creatively benefit from the guide For example creating useful tricks to aid you during a challenging test or quiz Chemistry may seem difficult but there are useful study methods O Level Chemistry Questions and Answers PDF Arshad Iqbal, The O Level Chemistry Quiz Questions and Answers PDF IGCSE GCSE Chemistry Competitive Exam Questions Chapter 1 14 Practice Tests Class 9 10 Chemistry Textbook Questions for Beginners includes revision guide for problem solving with hundreds of solved questions O Level Chemistry Questions and Answers PDF book covers basic concepts analytical and practical assessment tests O Level Chemistry Quiz PDF book helps to practice test questions from exam prep notes The O Level Chemistry Quiz Questions and Answers PDF eBook includes revision guide with verbal quantitative and analytical past papers solved tests O Level Chemistry Questions and Answers PDF Free download chapter 1 a book covers solved common questions and answers on chapters Acids and bases chemical bonding and structure chemical formulae and equations electricity electricity and chemicals elements compounds mixtures energy from chemicals experimental chemistry methods of purification particles of matter redox reactions salts and identification of ions and gases speed of reaction and structure of atom tests for school and college revision guide Chemistry Interview Questions and Answers PDF Download free eBook s sample covers beginner s solved questions textbook s study notes to practice online tests The IGCSE GCSE Chemistry Interview Questions Chapter 1 14 PDF book includes high school question papers to review practice tests for exams O Level Chemistry Practice Tests a textbook s revision guide with chapters tests for IGCSE NEET MCAT GRE GMAT SAT ACT competitive exam O Level Chemistry Questions Bank Chapter 1 14 PDF book covers problem solving exam tests from

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