

COMPOSITION OF HYDRATES

Topic: Hydrates

A hydrate is an ionic compound with water molecules loosely bonded to its crystal structure. The number of water molecules bonded to each unit of the salt. For example, the formula $\text{CuSO}_4 \cdot x\text{H}_2\text{O}$ indicates that there are x water molecules for every one formula unit of CuSO_4 . Answer the questions below.

1. What percentage of water is found in $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$?

$$C_{16} = 64$$

$$S_{32} = 32$$

$$O_{64} = 64$$

$$H_{10} = 10 \times 1 = 10$$

$$250 \text{ g CuSO}_4 = \frac{160}{250} \times 100 = 64\%$$

$$36\%$$

2. What percentage of water is found in $\text{FeSO}_4 \cdot x\text{H}_2\text{O}$?

$$\text{Fe}_{56} \quad 2S_{64} = 2 \times 32 = 64$$

$$O_{112} \quad 4H_{16} = 4 \times 1 = 16$$

$$H_{16} \quad 16 \times 1 = 16$$

$$\frac{160}{256} \times 100 = 62.5\%$$

$$37.5\%$$

3. A 5.0 gram sample of a hydrate of BaCl_2 was heated, resulting in 3.0 g of the anhydrous salt remaining. What percentage of water was in the hydrate?

$$\text{Ba}_{137} \quad 2Cl_{71} = 2 \times 35.5 = 71$$

$$\text{Water} = 2.0 \text{ g}$$

$$\text{Water} = 2.0 \text{ g} \times 100 = 40\%$$

$$\frac{2.0}{5.0} \times 100 = 40\%$$

4. A 2.5 g sample of a hydrate of CuSO_4 was heated, and only 1.7 g of the anhydrous salt remained. What percentage of water was in the hydrate?

$$\frac{0.8}{2.5} \times 100 = 32\%$$

$$\frac{0.8}{2.5} \times 100 = 32\%$$

$$\frac{0.8}{2.5} \times 100 = 32\%$$

5. A 3.0 gram sample of $\text{Na}_2\text{CO}_3 \cdot x\text{H}_2\text{O}$ is heated to constant weight. How much anhydrous salt remains?

$$2Na_{46} \quad C_{12} \quad O_{36} \quad 2H_{16} \quad 2O_{16} = 2 \times 16 = 32$$

$$\frac{186}{258} \times 100 = 72\%$$

$$\frac{186}{258} \times 100 = 72\%$$

$$\frac{186}{258} \times 100 = 72\%$$

$$28\%$$

6. A 5.0 gram sample of $\text{CuSO}_4 \cdot x\text{H}_2\text{O}$ was heated, and 3.0 g of the anhydrous salt remained. Determine the value of x .

$$C_{16} \quad S_{32} \quad O_{64} \quad 2H_{16} \quad 2O_{16} = 2 \times 16 = 32$$

$$2H_{16} \quad 2O_{16} = 2 \times 16 = 32$$

$$1 \times 18 = 18$$

$$\frac{160}{250} \times 100 = 64\%$$

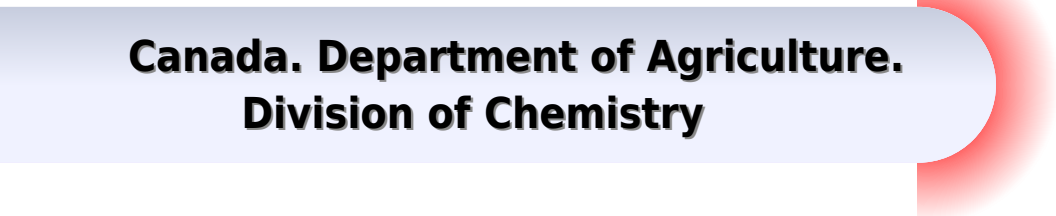
$$\frac{160}{250} \times 100 = 64\%$$

$$C_{16} \quad S_{32} \quad O_{64} \quad 2H_{16} \quad 2O_{16} = 2 \times 16 = 32$$

$$C_{16} \quad S_{32} \quad O_{64} \quad 2H_{16} \quad 2O_{16} = 2 \times 16 = 32$$

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