THE MOLE AND AVOGADRO'S NUMBER

Name _____

One male of a substance contains Avogadro's Number (6.02 x 1023) of molecules

How many molecules are in the quantities below?

- 1. 2.0 moles x = 20 mg (5.02 x 15 3 molecules) = 1.2 x 10 4 molecules
- 2. 1.5 moles x = 1.5 mbl × (6.02 × 13 3 molecules) = 9.0 × 10 molecules
- 3. 0.75 mole x = 0.75 m/ol x (6.02 x 133 molecular) = 4.5 x 10 molecular
- 4. 15 moles (6.02×152 molecula) = 9.0×154 molecules
- 5. 0.35 mole $\times (6.02 \times 10^{23} \text{ noleu. B.}) = 2.1 \times 10^{23} \text{ moleu. leu}$

How many moles are in the number of molecules below?

- 1. 6.02 × 1023 × rnde= = 6.07 × 13 3 moleculas (1000 (6.02 × 10) = 1.0 mol.
- 2. 1.204 x 1024 × moles: 1.204 x 134 molfale (1 mol (6,02 x 133 molfale) = 2.0 moles
- 3. 1.5 x 10°0 (0.00024 moles) (0.00024 moles (1.5 x 10° molegales (1.02 x 13° molegales) = 2.5 x 10° moles
- 4. 3.4 × 1026

 × moles: 3.4 × 15 moley les (1 mol 602 × 10° moleules) = 565 moles
- 5. 7.5 x 1010 × moles = 7.5 x 15 molearles (1 mol (6.02 x 13 molearles) = 1.2 x 10 moles

Chemistry If8766 Answer And Work

Lauren Gardner

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