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Chapter 5 – Chemical Reactions and Equations

- 5.1 (a) single-displacement reaction; (b) anhydrous; (c) molecular equation; (d) decomposition reaction; (e) balanced equation; (f) reactant; (g) spectator ion; (h) combustion reaction; (i) precipitate
- 5.2 (a) combination reaction; (b) activity series; (c) chemical equation; (d) product; (e) ionic equation; (f) neutralization reaction; (g) precipitation reaction; (h) net ionic equation; (i) double-displacement reaction
- 5.3 Substances that are formed as a result of the reaction are the products. This reaction produces solid aluminum oxide, $\text{Al}_2\text{O}_3(\text{s})$. You should pay attention to the substance formed *and* the state of the material if it is given. The reactants are the substances used to form the products. The reactants are aluminum metal, $\text{Al}(\text{s})$, and oxygen gas, $\text{O}_2(\text{g})$ (i.e. "Aluminum metal reacts with oxygen gas").
- 5.4 Substances that are formed as a result of the reaction are the products. Nitrogen gas and solid iodine, $\text{N}_2(\text{g})$ and $\text{I}_2(\text{s})$, and energy are the reaction products (i.e. "Nitrogen gas and solid iodine *form* in an explosive reaction"). You should pay attention to the substance formed and the state of the material if it is given. The reactants are the substances used to form the products. The reactant is nitrogen triiodide, NI_3 (i.e. "when nitrogen triiodide is *detonated*").
- 5.5 Image A represents the reactants because it contains both fluorine and xenon. Images B-D show product molecules, but only image C represents the products. In a chemical reaction, mass must be conserved. In image A, there are two xenon atoms and eight fluorine atoms. In order for mass to be conserved, the same number of each type of atom must be present in the products.
- 5.6 Make sure you look carefully at the images to distinguish between the chemical formulas for ozone (three red spheres) and oxygen (two red spheres). Image C represents the reactants because it contains both ozone and carbon monoxide. The products are represented by image D. Images A and B represent mixtures of reactants and products.
- 5.7 Mass must be conserved in a chemical reaction. In order for mass to be conserved, the same number of each type of atom must be present in both the products and reactants. The image is not accurate because mass is not conserved. To fix the image, you first need to count the number of each type of atom. Here's a summary of the atoms present in the reactants and products:

Atom	Reactants	Products
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Chemistry for Western Australia Cherie Lewis, Peter Lewis, 2008 Chemistry for Stage 2 Units 2A and 2B Provided by publisher

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