

Name **KEY** Test Date Period **CP Chemistry – Final Exam Review***Show your work and round using significant figure rules for all the problems.*

1. What is chemistry?

- **Chemistry is the study of matter and the changes it undergoes.**

2. Write the following in scientific notation:

a. .005 **5×10^{-3}**

c. 53099000 **5.3099×10^7**

b. 5050 **5.05×10^3**

d. 0.00053030 **5.3030×10^{-4}**

3. Write the following in standard notation:

a. 1.5×10^3 **1500**

c. 4×10^{-5} **0.00004**

b. 3.75×10^{-2} **0.0375**

d. 9.12×10^7 **91200000**

4. How many significant figures are in the following:

a. 0.02 **1**

c. 501 **3**

b. 501.0 **4**

d. 0.020 **2**

5. Calculate and report in correct significant figures:

a. $52.568 \text{ cm} + 689.45 \text{ cm} + 55.5694 \text{ cm} =$ **797.59 cm**

b. $8426.200 \text{ g} \div 254 \text{ mol} =$ **33.2 g/mol**

6. You measure a Calcium cube to have 2.01 cm sides and a mass of 11.6 grams:

a. What is the density as you calculate it?

$$V = s^3 = (2.01 \text{ cm})^3 = 8.12 \text{ cm}^3$$

$$d = \frac{m}{V} = \frac{11.6 \text{ g}}{8.12 \text{ cm}^3} = 1.43 \text{ g/cm}^3$$

b. If calcium's accepted density is 1.54 g/cm³, what is your percent error?

$$\% \text{error} = \frac{|1.54 \text{ g/cm}^3 - 1.43 \text{ g/cm}^3|}{1.54 \text{ g/cm}^3} \times 100 = 7.14 \%$$

7. Perform the following conversions:

a. 53.1 mm to m

$$53.1 \text{ mm} \times \frac{1 \text{ m}}{10^3 \text{ mm}} = 0.0531 \text{ m}$$

b. 18.4 kg to dag

$$18.4 \text{ kg} \times \frac{1 \times 10^{-1} \text{ dag}}{1 \times 10^{-3} \text{ kg}} = 1840 \text{ dag}$$

c. 93.6 mi to km (1 mi = 1.6 km)

$$93.6 \text{ mi} \times \frac{1.6 \text{ km}}{1 \text{ mi}} = 150. \text{ km}$$

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Daniel C. Tofan



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1. Understanding the eBook Chemistry Final Study Guide Answers
 - The Rise of Digital Reading Chemistry Final Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Final Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Final Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Final Study Guide Answers
 - Personalized Recommendations
 - Chemistry Final Study Guide Answers User Reviews and Ratings
 - Chemistry Final Study Guide Answers and Bestseller Lists
5. Accessing Chemistry Final Study Guide Answers Free and Paid eBooks
 - Chemistry Final Study Guide Answers Public Domain eBooks
 - Chemistry Final Study Guide Answers eBook Subscription Services
 - Chemistry Final Study Guide Answers Budget-Friendly Options
6. Navigating Chemistry Final Study Guide Answers eBook Formats

- ePub, PDF, MOBI, and More
- Chemistry Final Study Guide Answers Compatibility with Devices
- Chemistry Final Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Final Study Guide Answers
 - Highlighting and Note-Taking Chemistry Final Study Guide Answers
 - Interactive Elements Chemistry Final Study Guide Answers
- 8. Staying Engaged with Chemistry Final Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Final Study Guide Answers
- 9. Balancing eBooks and Physical Books Chemistry Final Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Final Study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Final Study Guide Answers
 - Setting Reading Goals Chemistry Final Study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Final Study Guide Answers
 - Fact-Checking eBook Content of Chemistry Final Study Guide Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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