

## SOLUTION

1.	Solution: A homogeneous mixture of two
	or more substances is known as solution.
2.	Solute: The substance present in smaller amount in a solution is called solute.
3.	Solvent: The substance present in larger

- Solvent: The substance present in larger amount in a solution is called solvent.
- Binary solution: Theose solutions which contain two components are called binary solutions, e.g., salt solution, benzene and toluene.
- Ternary solutions: Those solutions which contain three components are called ternary solutions, e.g., ethanol + water + acetic acid.
- Aqueous solution: When solute is dissolved in water, it is called aqueous solution, e.g., sugar solution, ethanol in water.
- Non-aqueous solution: When solute is dissolved in solvent other than water, it is called non-aqueous solution, e.g., iodine dissolved in alcohol (Tincture of iodine).
- 8. Weight Fraction = weight of solution
- Weight percentage

Weight of solute present in 100 grams of solution is called weight percentage

Volume of solute present in 100 ml of the solution is called volume percentage

- 11. Volume Fraction = Volume of solution
- 12. Molarity (M)

(or) 
$$M = \frac{W}{GMW} \times \frac{1}{V(lit)}$$

13. Molarity = No of moles of solute

14. Molarity =  $\frac{\text{Weight of solute}}{\text{GMWt of solute}} \times \frac{1000}{\text{Vin ml}}$ 

(or) 
$$M = \frac{W}{GMW} \times \frac{1000}{V(ml)}$$

- 15. Weight of solute =  $\frac{M \times GMW \times Vinml}{1000}$
- 16. Molarity for dilute solutions:  $M_iV_i = M_iV_j$
- 17. For neutralisation reactions:  $\frac{M_1V_1}{n_1} = \frac{M_2V_2}{n_2}$
- 18. For mixed solution molarity

$$= \frac{M_1V_1 + M_2V_2 + .....}{V_1 + V_2 + .....}$$

# **Chemistry Dimensions 2 Solutions**

G. V. Kozlov, I. V. Doblin, Gennady E. Zaikov

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solutions as one possibility The challenge was taken up by Marie Claire Bellissent Funel and George Neilson I am sure that all the participants would wish to congratulate our two colleagues for putting together an outstanding programme of lectures round tables and poster session The theory which underlies the subject was covered by four leading authorities I P Hansen Paris set out the general framework in terms of the statistical mechanics of bulk and surface properties H L Friedman Stony Brook focused attention on ionic liquids at equilibrium and I B Hubbard considered non equilibrium properties such as the electrical conductivity and ionic friction coefficients Finally the basic theory of polyelectrolytes treated as charged linear polymers in aqueous solution was presented by J M Victor Paris **Physical Chemistry for Students of Medicine** The Fractal Physical Chemistry of Polymer Solutions and Melts G. V. Kozlov, I. V. Alexander Findlay, 1924 Doblin, Gennady E. Zaikov, 2013-12-12 This book provides an important structural analysis of polymer solutions and melts using fractal analysis The book covers the theoretical fundamentals of macromolecules fractal analysis It then goes on to discuss the fractal physics of polymer solutions and the fractal physics of melts The intended audience of the book includes specialists in chemistry and physics of polymer synthesis and those in the field of polymers and polymer composites Geopolymer, Green Chemistry and Sustainable Development Solutions Joseph Davidovits, 2005 processing

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