

EQUILIBRIUM CHEAT SHEET

Chemical equilibrium, dynamic equilibrium, when the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium reactions are represented by this symbol: $\rightleftharpoons$

What is Equilibrium?

Equilibrium is a constant concentration of reactants and products.

When a reversible reaction takes place in a closed system, the forward and reverse reactions occur at the same rate. The system is called dynamic equilibrium if the concentrations of reactants and products do not change.

It is interesting to think of equilibrium as a constant state, but the forward and reverse reactions are still occurring. The system is in a constant state of change.

Equilibrium Constant Formed
 $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$

Equilibrium Constant
 $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$

Example:
 $2 H_2 + O_2 \rightleftharpoons 2 H_2O$

Forward Reaction:
 $2 H_2 + O_2 \rightarrow 2 H_2O$

Reverse Reaction:
 $2 H_2O \rightarrow 2 H_2 + O_2$



Equilibrium Constant, K_c

The equilibrium constant, K_c , is a measure of the extent to which a reaction proceeds. It is calculated as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.

$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$

Equilibrium Constant Equation:

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

The equilibrium constant, K_c , is a measure of the extent to which a reaction proceeds. It is calculated as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.

Equilibrium Constant (Kc) Writing	
Equilibrium Constant Value	Meaning
$K_c > 1$	Concentration of products is greater than concentration of reactants
$K_c < 1$	Concentration of reactants is greater than concentration of products
$K_c = 1$	Reactants and products are present in equal amounts

Reaction Quotient, Q

The reaction quotient, Q , is a measure of the extent to which a reaction proceeds. It is calculated as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.

$Q = \frac{[C]^c [D]^d}{[A]^a [B]^b}$

The reaction quotient, Q , is a measure of the extent to which a reaction proceeds. It is calculated as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.

Reaction Quotient vs. Equilibrium Constant	
Equilibrium Constant (Kc) Value	Meaning
$Q < K_c$	Reaction is not at equilibrium. Forward reaction is favored.
$Q > K_c$	Reaction is not at equilibrium. Reverse reaction is favored.
$Q = K_c$	Reaction is at equilibrium.

Le Chatelier's Principle: The effect of changing conditions on equilibrium		
Condition	Effect of Change	Example
Concentration	- Increasing concentration of reactants shifts equilibrium to the right. - Increasing concentration of products shifts equilibrium to the left.	$2 H_2 + O_2 \rightleftharpoons 2 H_2O$ Increasing concentration of H_2 or O_2 shifts equilibrium to the right, increasing concentration of H_2O .
Pressure	The equilibrium shifts to the side with fewer moles of gas. - Increasing pressure shifts equilibrium to the side with fewer moles of gas. - Decreasing pressure shifts equilibrium to the side with more moles of gas.	$2 H_2 + O_2 \rightleftharpoons 2 H_2O$ Increasing pressure shifts equilibrium to the right, increasing concentration of H_2O .
Temperature	The equilibrium shifts to the side that absorbs heat. - Increasing temperature shifts equilibrium to the side that absorbs heat. - Decreasing temperature shifts equilibrium to the side that releases heat.	$2 H_2 + O_2 \rightleftharpoons 2 H_2O$ Increasing temperature shifts equilibrium to the left, decreasing concentration of H_2O .

Chemistry Chemical Equilibrium Study Guide Answers

Hussin A.Rothana



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