Organic Chemistry - Some Basic Principles and Techniques

Question 8.1:

What are hybridisation states of each carbon atom in the following compounds? $CH_2 = C = O$, $CH_3CH = CH_2$, $(CH_3)_2CO$, $CH_2 = CHCN$, C_6H_6

Answer 8.1:

- (i) $\dot{C}H_2 = \dot{C} = 0$ C-1 is sp^2 hybridised. C-2 is sp hybridised.
- (ii) $\overset{\circ}{C}H_3 \overset{\circ}{C}H = \overset{\circ}{C}H_2$ C-1 is sp^3 hybridised. C-2 is sp^2 hybridised. C-3 is sp^2 hybridised.

- (iv) $\overset{\cdot}{C}H_{z} = \overset{\cdot}{C}H \overset{\cdot}{C} = N$ C-2 is sp^{2} hybridised. C-3 is sp hybridised.
- (v) C_sH_s
 All the 6 carbon atoms in benzene are sp² hybridised.

Question 8.2:

Indicate the σ and π bonds in the following molecules: C_6H_6 , C_6H_{12} , CH_2Cl_2 , $CH_2=C=CH_2$, CH_3NO_2 , $HCONHCH_3$

Answer 8.2:

- (i) C°H°
- There are six C-C sigma (σ_{c-c}) bonds, six C-H sigma (σ_{c-H}) bonds and three C=C pi (π_{c-c}) resonating bonds in this compound.

There are six C–C sigma (σ_{c-c}) bonds and twelve C–H sigma (σ_{c-c}) bonds in this compound.

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