

Stoichiometry

Section 11.1 Defining Stoichiometry

pages 368–372

Practice Problems

pages 371–372

1. Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that the law of conservation of mass is observed.



1 molecule N_2 + 3 molecules

$\text{H}_2 \rightarrow$ 2 molecules NH_3

1 mole N_2 + 3 moles $\text{H}_2 \rightarrow$ 2 moles NH_3

Mass:

N_2 : (2 mol)(14.007 g/mol) = 28.014 g

3H_2 : (6 mol)(1.008 g/mol) = 6.048 g

2NH_3 : (2 mol)(14.007 g/mol) +
(6 mol)(1.008 g/mol) = 34.062 g

28.014 g N_2 + 6.048 g $\text{H}_2 \rightarrow$ 34.062 g NH_3

34.062 g reactants = 34.062 g products



1 molecule HCl + 1 formula unit $\text{KOH} \rightarrow$

1 formula unit KCl + 1 molecule H_2O

1 mole HCl + 1 mole $\text{KOH} \rightarrow$

1 mole KCl + 1 mole H_2O

Mass:

HCl : (1 mol)(1.008 g/mol) +
(1 mol)(35.453 g/mol) = 36.461 g

KOH : (1 mol)(39.098 g/mol) +
(1 mol)(15.999 g/mol) + (1 mol)(1.008 g/mol)
= 56.105 g

KCl : (1 mol)(39.098 g/mol) +
(1 mol)(35.453 g/mol) = 74.551 g

H_2O : (2 mol)(1.008 g/mol) +
(1 mol)(15.999 g/mol) = 18.015 g

36.461 g HCl + 56.105 g $\text{KOH} \rightarrow$
74.551 g KCl + 18.015 g H_2O

92.566 g reactants = 92.566 g products



2 atoms Mg + 1 molecule

$\text{O}_2 \rightarrow$ 2 formula units MgO

2 moles Mg + 1 mole $\text{O}_2 \rightarrow$ 2 moles MgO

Mass:

2Mg : (2 mol)(24.305 g/mol) = 48.610 g

O_2 : (2 mol)(15.999 g/mol) = 31.998 g

2MgO : (2 mol)(24.305 g/mol) +
(2 mol)(15.999 g/mol) = 80.608 g

48.610 g Mg + 31.998 g $\text{O}_2 \rightarrow$ 80.608 g MgO

80.608 g reactants = 80.608 g products

2. **Challenge** For each of the following, balance the chemical equation; interpret the equation in terms of particles, moles, and mass; and show that the law of conservation of mass is observed.



2 atoms Na + 2 molecules $\text{H}_2\text{O} \rightarrow$

2 formula units NaOH + 1 molecule H_2

2 mol Na + 2 mol $\text{H}_2\text{O} \rightarrow$

2 mol NaOH + 1 mol H_2

Mass:

2Na : (2 mol)(22.990 g/mol) = 45.980 g

$2\text{H}_2\text{O}$: (4 mol)(1.008 g/mol) +
(2 mol)(15.999 g/mol) = 36.030 g

2NaOH : (2 mol)(22.990 g/mol) +
(2 mol)(15.999 g/mol) + (2 mol)(1.008 g/mol)
= 79.994 g

H_2 : (2 mol)(1.008 g/mol) = 2.016 g

45.980 g Na + 36.030 g $\text{H}_2\text{O} \rightarrow$
79.994 g NaOH + 2.016 g H_2

82.01 g reactants = 82.01 g products

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James E. Brady, Fred Senese



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Table of Contents Chemistry Chapter 11 Stoichiometry Study Guide Answers

1. Understanding the eBook Chemistry Chapter 11 Stoichiometry Study Guide Answers
 - The Rise of Digital Reading Chemistry Chapter 11 Stoichiometry Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 11 Stoichiometry Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Chemistry Chapter 11 Stoichiometry Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 11 Stoichiometry Study Guide Answers
 - Personalized Recommendations
 - Chemistry Chapter 11 Stoichiometry Study Guide Answers User Reviews and Ratings
 - Chemistry Chapter 11 Stoichiometry Study Guide Answers and Bestseller Lists
5. Accessing Chemistry Chapter 11 Stoichiometry Study Guide Answers Free and Paid eBooks
 - Chemistry Chapter 11 Stoichiometry Study Guide Answers Public Domain eBooks
 - Chemistry Chapter 11 Stoichiometry Study Guide Answers eBook Subscription Services
 - Chemistry Chapter 11 Stoichiometry Study Guide Answers Budget-Friendly Options
6. Navigating Chemistry Chapter 11 Stoichiometry Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 11 Stoichiometry Study Guide Answers Compatibility with Devices
 - Chemistry Chapter 11 Stoichiometry Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 11 Stoichiometry Study Guide Answers
 - Highlighting and Note-Taking Chemistry Chapter 11 Stoichiometry Study Guide Answers
 - Interactive Elements Chemistry Chapter 11 Stoichiometry Study Guide Answers

8. Staying Engaged with Chemistry Chapter 11 Stoichiometry Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 11 Stoichiometry Study Guide Answers
9. Balancing eBooks and Physical Books Chemistry Chapter 11 Stoichiometry Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 11 Stoichiometry Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 11 Stoichiometry Study Guide Answers
 - Setting Reading Goals Chemistry Chapter 11 Stoichiometry Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 11 Stoichiometry Study Guide Answers
 - Fact-Checking eBook Content of Chemistry Chapter 11 Stoichiometry Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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