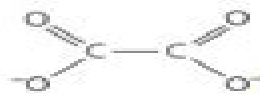




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21: Transition metals

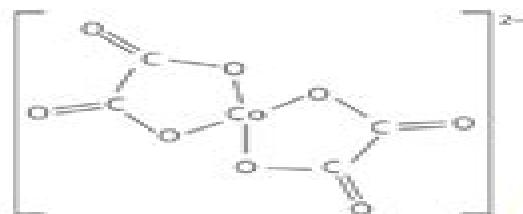
- 1 (a) At least one ligand donates an electron pair [1]; to a transition metal ion to form a coordinate bond [1]
- (b) CH_4 [1]; no lone pair available to form a coordinate bond [1]
- (c) (i) +3 [1]
 (ii) 4 [1]
 (iii) Tetrahedral [1]; square planar [1]
- 2 (a) (i) Electrons are promoted from a lower energy state to a higher energy state [1]; the energy difference, ΔE , between these two energy states determines the frequency of visible light absorbed [1]; pink light is transmitted because it is not absorbed [1]
 (ii) ΔE is the energy difference between the lower and upper energy states [1]; h is the Planck constant [1]; ν is the frequency of the radiation emitted [1]
 (iii) $[\text{Co}(\text{H}_2\text{O})_6]^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightarrow [\text{CoCl}_4]^{2-}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$ [1] for all chemical species; [1] for balancing
 (iv) Tetrahedral [1]; chloride ions are larger than oxygen atoms in water molecules [1]; more space needed/more repulsion between ligands [1]
 (v)



[1]; two lone pairs available, one on each carboxylate group [1]

- (b) (i) $[\text{Co}(\text{H}_2\text{O})_6]^{2+}(\text{aq}) + 2\text{C}_2\text{O}_4^{2-}(\text{aq}) \rightarrow [\text{Co}(\text{C}_2\text{O}_4)_2(\text{H}_2\text{O})_2]^{2-}(\text{aq}) + 4\text{H}_2\text{O}(\text{l})$ [2]
 coordination number = 6 [1]; octahedral [1]

(ii)



[2]

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**T. W. Graham Solomons, Craig B.
Fryhle, Scott A. Snyder**



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