

Acids & Bases Notes

Lesson 1

Recall that acids begin with H and bases end with OH. There are exceptions, commonly ammonia (NH₃) is a base.

Remember that there are strong and weak acids and bases. List the strong acids:

HCl, HBr, HI, H₂SO₄, HNO₃, HClO₄

Know (use back of C, but Hydrochloric, Hydrobromic, Hydroiodic, Sulfuric, Nitric, and perchloric acid)

List the strong bases:

LiOH, NaOH, KOH, RbOH, CsOH, FrOH Hydroxides of the alkali metals (group 1 metals)

Ca(OH)₂, Sr(OH)₂, and Ba(OH)₂

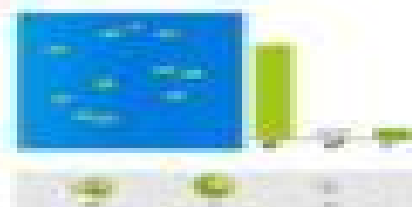
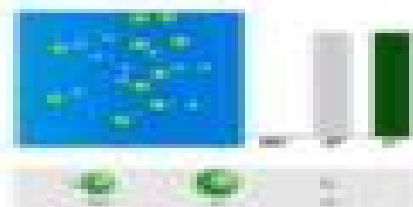
Pb(OH)₂, Pb(OH)₃

Identify Pb(OH)₂ as an acid or base. Is it strong or weak? weak base



What makes an acid/base "strong" or "weak"?

Strong acids and bases completely dissociate (break apart) in water. Weak acids and bases only partly dissociate.



Practice Problems

Draw an image showing HCl dissolving in water.

Draw an image of NaOH dissolving in water.



Acids form H⁺ in water. Bases form OH⁻ in water.



H⁺ combined with H₂O to form H₃O⁺. H₃O⁺ is called hydronium.



Strong acids form a lot of H₃O⁺ in water. Weak acids don't form much H₃O⁺.



Chemistry Acids And Bases Guided Answers

Michael Hoelscher



Chemistry Acids And Bases Guided Answers:

This book delves into Chemistry Acids And Bases Guided Answers. Chemistry Acids And Bases Guided Answers is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Chemistry Acids And Bases Guided Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Chemistry Acids And Bases Guided Answers
 - Chapter 2: Essential Elements of Chemistry Acids And Bases Guided Answers
 - Chapter 3: Chemistry Acids And Bases Guided Answers in Everyday Life
 - Chapter 4: Chemistry Acids And Bases Guided Answers in Specific Contexts
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