CHEMICAL QUANTITIES STUDY GUIDE

Atomic Mass

- The mass of a single atom of an element (amu)
- Mass of one mole of the atom = atomic mass x amu
 - 1 amu = 1.66 x 10⁻²⁴ g

Avogadro's Number

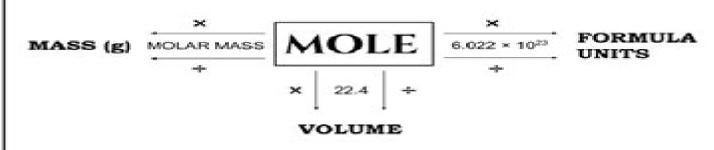
- 6.022 x 10²⁰
- Six hundred 2 sextillion two hundred quintillion
- Represents the number of atoms that are in 1 mole of a substance

Molar Mass

- The sum of the total mass in grams of all the atoms that make up a mole of a particular molecule
- Molar mass of an element = number underneath name on periodic table
 - For a compound, calculate molar mass of each element and add

Mole Conversions

Moles = mass/molar mass



Chemical Quantities Study Guide

National Environmental Health Association (NEHA)

Chemical Quantities Study Guide:

Study Guide/Selected Solutions Manual Julie R. Frentrup,2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

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