

CHEMFILE MINI-GUIDE TO PROBLEM SOLVING

Gas Laws

Chemists found that there were relationships among temperature, volume, pressure, and quantity of a gas that could be described mathematically. This chapter deals with Boyle's law, Charles's law, Gay-Lussac's law, the combined gas law, and Dalton's law of partial pressures. These laws have one condition in common. They all assume that the molar amount of gas does not change. In other words, these laws work correctly only when no additional gas enters a system and when no gas leaks out of it. Remember also that a law describes a fact of nature. Gases do not "obey" laws. The law does not dictate the behavior of the gas. Rather, each gas law describes a certain behavior of gas that occurs if conditions are right.

BOYLE'S LAW

Robert Boyle, a British chemist who lived from 1627 to 1691 formulated the first gas law, now known as Boyle's law. This law describes the relationship between the pressure and volume of a sample of gas confined in a container. Boyle found that gases compress, much like a spring, when the pressure on the gas is increased. He also found that they "spring back" when the pressure is lowered. By "springing back" he meant that the volume increases when pressure is lowered. It's important to note that Boyle's law is true only if the temperature of the gas does not change and no additional gas is added to the container or leaks out of the container.

Boyle's law states that *the volume and pressure of a sample of gas are inversely proportional to each other at constant temperature*. This statement can be expressed as follows.

Proportionality symbol. It means "is proportional to."

Proportionality constant

$$V \propto \frac{1}{P} \quad \text{and} \quad PV = k \quad \text{or} \quad V = k \frac{1}{P}$$

According to Boyle's law, when the pressure on a gas is *increased*, the volume of the gas *decreases*. For example, if the pressure is doubled, the volume decreases by half. If the volume quadruples, the pressure decreases to one-fourth of its original value.

Chemfile Mini Guide To Gas Laws

AW Rasmussen



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Holt Chemistry File, 1998 This reference is a must for students who need extra help reteaching or extra practice The guide moves students through the same concepts as the text but at a slower pace More descriptive detail along with visual algorithms provides a more structured approach Each chapter closes with a large bank of practice problems Book jacket

Expanding on the Gas Laws, 2009-01-01 Inquiries in Science Chemistry Series Expanding on the Gas Laws Teacher's Guide *Ideal Gas Law 51 Success Secrets - 51 Most Asked Questions on Ideal Gas Law - What You Need to Know* Philip Hicks, 2014-10-07 It's a brand new Ideal gas law world There has never been a Ideal gas law Guide like this It contains 51 answers much more than you can imagine comprehensive answers and extensive details and references with insights that have never before been offered in print Get the information you need fast This all embracing guide offers a thorough view of key knowledge and detailed insight This Guide introduces what you want to know about Ideal gas law A quick look inside of some of the subjects covered Atmospheric thermodynamics Overview Thermodynamic instruments Thermodynamic meters Glossary of engineering I Idealization Limits on use Perfect gas Stoichiometry Water vapor Water vapor and dry air density calculations at 0 C Equipartition theorem Perfection Physics and chemistry Glossary of chemistry terms U Fusion energy 1960s Timeline of low temperature technology 19th century Gas Avogadro's law Hot air balloon List of multiple discoveries 17th century Amount of substance Equation of state Overview Explosive Volume of products of explosion Aerodynamics Conservation laws Van der Waals equation Validity Equipartition of energy Gas Physical characteristics Gas meter Flow measurement calculations Mass flow sensor Chamber pressure Importance in Firearm Maintenance Weather forecasting How models create forecasts Timeline of hydrogen technologies 1800s Pressure Pressure of an ideal gas Compressible fluid One Dimensional Flow Diffusion Elementary theory of diffusion coefficient in gases Water vapour Water vapor and Density of airdry air density calculations at 0 C Ideal gas law Numerical weather prediction Computation Gay Lussac's law Pressure temperature law Hydrostatic equilibrium Astrophysics History of thermodynamics Birth of thermodynamics as science and much more **Gases** Renée G. Ford, 1962 *Gas Laws* Lifeliqe, 2019 This lesson plan covers Boyle's law Charles law Gay Lussac's law and Avogadro's law The laws are applied to make various calculations concerning constant temperature pressure volume and moles *The Gas Laws* Irene Cesa, 2003-01-01 *The Laws of Gases* Robert Boyle, Emile Hilaire Amagat, 1899 **Gas Laws** Source Wikipedia, 2013-09 Please note that the content of this book primarily consists of articles available from Wikipedia or other free sources online Pages 24 Chapters Acentric factor Amagat's law Avogadro's law Boyle's law Charles's law Combined gas law Compressibility factor Dalton's law Gay Lussac's law Graham's law Henry's law Magic number chemistry Partial pressure Psychrometric constant Redlich Kwong equation of state Van der Waals constants data page Van der Waals equation Excerpt The van der Waals equation is an equation of state for a fluid composed of particles that have a non zero volume and a pairwise attractive inter particle force such as the van der Waals force It was derived in

1873 by Johannes Diderik van der Waals who received the Nobel prize in 1910 for his work on the equation of state for gases and liquids. The equation is based on a modification of the ideal gas law and approximates the behavior of real fluids taking into account the nonzero size of molecules and the attraction between them. The van der Waals isotherms the model correctly predicts a mostly incompressible liquid phase but the oscillations in the phase transition zone do not fit experimental data. The equation uses the following state variables: the pressure of the fluid p , total volume of the container containing the fluid V , number of moles n , and absolute temperature of the system T . One form of the equation is where v is the volume of the container shared between each particle, u is the velocity of a particle, N is the total number of particles, and k_B is Boltzmann's constant given by the universal gas constant R and Avogadro's constant N_A . Extra parameters are introduced: a is a measure for the attraction between the particles and b is the average volume excluded from v by a particle. The equation can be cast into the better known form where a is a measure of the attraction between the particles, b is the volume excluded by a mole of particles. A careful distinction.

Introductory Chemistry, Module Five : the Gas Laws Hoewe, Barbara, Jalil, Shahid, 1976

The Laws of Gases: Memoirs Carl Barus, Robert Boyle, Émile Hilaire Amagat, 2023-07-18

This scientific memoir covers the fundamental laws and properties of gases as understood by three pioneering researchers in the field. Through experiments and rigorous analysis, the authors explore topics such as temperature, pressure, heat, and chemical reactions, shedding light on the complex and fascinating dynamics of gas molecules. With detailed descriptions and illustrations, this book provides a comprehensive introduction to the subject for students and scholars alike. This work has been selected by scholars as being culturally important and is part of the knowledge base of civilization as we know it. This work is in the public domain in the United States of America and possibly other nations. Within the United States, you may freely copy and distribute this work as no entity individual or corporate has a copyright on the body of the work. Scholars believe and we concur that this work is important enough to be preserved, reproduced, and made generally available to the public. We appreciate your support of the preservation process and thank you for being an important part of keeping this knowledge alive and relevant.

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