

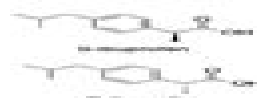
Introduction to Chem 2443:

Organic chemistry = the branch of science that deals generally with compounds of carbon regardless of their origin

- Synthesis of organic compounds can be done in a lab

Inorganic = not present in living systems/organisms

- Molecular structure determines function!
 - Proteins – (1) structural and (2) active/enzymes
 - Fold because of hydrophobic parts that don't want to interact with water
 - 3D shape is very important
 - Solid line = coming towards
 - Dashed lines = going away



Bonding and Hybridization:

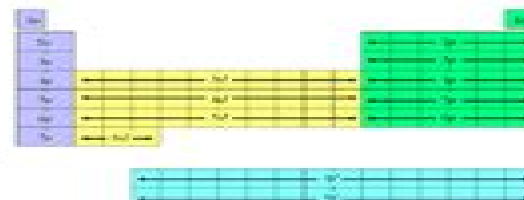
- Atomic identity is based on the number of protons
 - Isotopes = different number of neutrons, but doesn't change the identity (protons)
- Bottom left = atomic number (protons=electrons)
- Upper left = atomic mass (protons + neutrons)



- Formal charges indicate when an atom has more or less valence electrons than what it should normally have



*does not affect electronegativity



- Only valence electrons participate in chemical bonding
 - Sharing of V.E.s between atoms = covalent bonds
 - Covalent bonding is thermodynamically favored for stability
 - Hybrid orbitals are used for covalent bonding

2 "things" →

3 "things" →

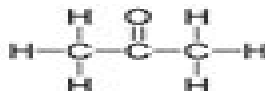
Types of hybridization	Geometry
sp hybridization	Linear
sp ² hybridization	Trigonal planar
sp ³ hybridization	Tetrahedral
dsp ² hybridization	Square planar
dsp ³ hybridization	Trigonal bipyramidal
d ² sp ³ hybridization	Octahedral
d ³ sp ³ hybridization	Pentagonal bipyramidal

"things" = lone pairs, single bond, double bond, triple bond

O always has 2 lone pairs, C has no lone pairs, N has one lone pair

O: sp²

C: sp²



N: sp³



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David R. Klein



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