

CHEM 151 – Quiz #4
Thermo, Redox, and Electrochem
 December 4, 2018

1. Consider this unbalanced reaction at 25 °C:

Zinc solid reacts with a solution containing nitrate ion, to form zinc ion and nitrogen monoxide gas.

The standard reduction potentials (E°) are:

Nitrogen half-reaction:	+ 0.957 V	red
Zinc half-reaction:	- 0.762 V	ox

- a. (2 pts) Using chemical formulas, write the unbalanced reaction given above.



- b. (5 pts) Balance the reaction in an acidic solution using the half-reaction method practiced in class (include phases in your final balanced reaction).



- c. (3 pts) i) assign oxidation numbers to all atoms in your final balanced reaction, ii) calculate the change of oxidation state for the two atoms that underwent change, and identify if each atom underwent oxidation or reduction, and iii) label the oxidizing and reducing agents (from the reactants only).

i) see above

ii) $\text{Zn} \xrightarrow{\text{iii)}} 0 \rightarrow +2$, oxidation $\Rightarrow \text{Zn(s)} = \text{reducing agent}$

$\text{N} \xrightarrow{\text{iii)}} +5 \rightarrow +2$, reduction $\Rightarrow \text{NO}_3^-(\text{aq}) = \text{oxidizing agent}$

Chem 151 Lab Manual Answer Key

Elizabeth Kean, Catherine Middlecamp



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1. Understanding the eBook Chem 151 Lab Manual Answer Key
 - The Rise of Digital Reading Chem 151 Lab Manual Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chem 151 Lab Manual Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chem 151 Lab Manual Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chem 151 Lab Manual Answer Key
 - Personalized Recommendations
 - Chem 151 Lab Manual Answer Key User Reviews and Ratings
 - Chem 151 Lab Manual Answer Key and Bestseller Lists
5. Accessing Chem 151 Lab Manual Answer Key Free and Paid eBooks
 - Chem 151 Lab Manual Answer Key Public Domain eBooks
 - Chem 151 Lab Manual Answer Key eBook Subscription Services
 - Chem 151 Lab Manual Answer Key Budget-Friendly Options

6. Navigating Chem 151 Lab Manual Answer Key eBook Formats
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 - Chem 151 Lab Manual Answer Key Compatibility with Devices
 - Chem 151 Lab Manual Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chem 151 Lab Manual Answer Key
 - Highlighting and Note-Taking Chem 151 Lab Manual Answer Key
 - Interactive Elements Chem 151 Lab Manual Answer Key
8. Staying Engaged with Chem 151 Lab Manual Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chem 151 Lab Manual Answer Key
9. Balancing eBooks and Physical Books Chem 151 Lab Manual Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chem 151 Lab Manual Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chem 151 Lab Manual Answer Key
 - Setting Reading Goals Chem 151 Lab Manual Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chem 151 Lab Manual Answer Key
 - Fact-Checking eBook Content of Chem 151 Lab Manual Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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