

Chemistry 101 L: General Chemistry I Laboratory

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Pierce College

Fall 2011 Sections 8077

Office Phone: (818)712-2618

DESCRIPTION: General chemistry I presents the principles and laws of chemistry as related to the structure of matter. Topics covered include a comparison of the states of matter; atomic structure and the periodic table; stoichiometry; thermochemistry and introductory thermodynamics; chemical bonding; solutions; solubility; acids and bases; introductory chemical equilibrium; phase changes; and an introduction to Molecular Orbital Theory. The laboratory work is intended to develop skills in observation, use of chemical glassware and equipment, making deductions from observations, analyzing results and communicating them in a written laboratory report.

PREREQUISITES: Chem 60 and Math 115 with a "C" grade or better or equivalent skill level demonstrated through the mathematics placement process. It is **strongly recommended** that you take English 28 before you attempt this class in order for you to have the writing skills necessary for this class.

CHEMISTRY 101: General Chemistry I Student Learning Outcomes (SLOs)

1. Conduct experiments and interpret data. This includes the ability to :
 - (a) perform experiments to answer questions and characterize unknowns;
 - (b) record, manipulate and evaluate the experimental data using graphs and equations as needed, to reach conclusions;
 - (c) correlate experimental results with the appropriate theory and
 - (d) use instrumentation and appropriate laboratory techniques to carry out experiments.
2. Solve mathematical problems specific to chemistry that require identifying key data (from lists, tables, experiments or graphs) and constructing correct formulas for unit conversions, ratios, and stoichiometry.
3. Obtain and analyze data from different sources (i.e., literature/information technology).
4. Apply the atomic and kinetic theories of matter to how atomic and molecular properties are responsible for chemical and physical behavior at the macroscopic level and its impact on the global environment.
5. Compare and contrast the details of ionic, covalent and intermolecular bonding. This includes describing how energy changes are related to temperature, motion at the atomic level, and changes in chemical bonding.
6. Practice safe and effective general laboratory skills, individually and in groups. This includes
 - (a) recognizing the limitations of physical measurements and the application of appropriate rules for significant figures;
 - (b) completing measurements accurately and precisely; and
 - (c) effectively working with peers in a collegial environment.

ATTENDANCE: The guidelines in the course catalog apply to this class: Any student who has unexcused absences equaling one week's worth of class time prior to census date will be excluded. **It is your responsibility to drop this class if necessary.** Three lateness's will count as an absence. Please contact the instructor if you know you are going to miss several days of class. Please read the attendance regulations printed in the *Schedule of Classes*.

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James Winfred Neckers



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