

1. Here are three isotopes of an element: $^{12}_6\text{C}$ $^{13}_6\text{C}$ $^{14}_6\text{C}$
 - a. The element is: **Carbon**
 - b. The number 6 refers to the: **Atomic Number**
 - c. The numbers 12, 13, and 14 refer to the **different mass numbers**
 - d. How many protons and neutrons are in the first isotope? **P: 6 N: 6**
 - e. How many protons and neutrons are in the second isotope? **P: 6 N: 7**
 - f. How many protons and neutrons are in the third isotope? **P: 6 N: 8**

2. Complete the following chart:

Isotope name	atomic #	mass #	# of protons	# of neutrons	# of electrons
Potassium-37	19	37	19	18	19
Oxygen-17	8	17	8	9	8
uranium-235	92	235	92	143	92
uranium-238	92	238	92	146	92
boron-10	5	10	5	5	5
boron-11	5	11	5	6	5

3. How do you calculate the mass number of an element?

Mass # = protons + neutrons

4. If an atom has 14 protons, 17 neutrons and 14 electrons, which expression identifies the element?
 - a. Silicon-14
 - b. Silicon-28
 - c. **Silicon-31**
 - d. Silicon-17
5. If an atom has 53 protons, 54 neutrons and 53 electrons, which expression identifies the element?
 - a. Iodine-53
 - b. Iodine-127
 - c. Iodine-54
 - d. **Iodine-107**

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Mark Steyvers



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