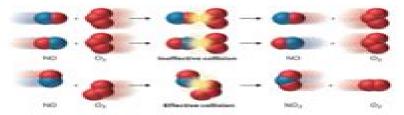
#### Collision Theory Summary

Main Idea: Particles must collide to react.

However, not all collisions are successful/effective. In other words, not all collisions result in a chemical reaction.

In order for reactants to be converted to products two requirements must be met:

- The collision must have sufficient energy to overcome the required activation energy of the reaction.
- The particles must collide at the correct orientation to allow a reaction to occur.



There are three main ways to increase successful/effective collisions:

- Increase the collision frequency. (Increase how often the particles collide.)
- Increase the collision energy.
- Increase correct collision orientation.

The rate of a chemical reaction is dependent on the rate of successful particle collisions.

Using the above ideas, we can describe five factors (nature of reactants, concentration, addition of a catalyst, exposed surface area, temperature) that affect the rate of a reaction.

Of the five factors which factors will affect each of the following variables:

Collision Frequency	Collision Energy	Collision Orientation
Particle speed affects how often the particles collide  Concentration The number of particles in a given space affects how often the particles collide  Nature of Reactants The state/phase of the reactants affects their movement and therefore, how often the particles collide  Surface Area The amount of exposure the reactant particles have to each other affects how often the particles collide	Temperature Particle speed affects the energy/force the particles collide with	Nature of Reactants  Different substances have particles with different sizes, shapes and complexity. Particles that are more complex and non-symmetrical will need to collide in a more specific way (ex. with one end or part of the particle) than simpler particles.  Catalysts  Catalysts  Catalysts provide an alternate way for the reaction to occur, often by orienting the reacting particles in such a way that successful collisions are more likely

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estimates of kinetic parameters have been included in this version of the book Finally kinetics requires knowledge that must be complemented and tested in the laboratory Therefore practical examples of reactions performed in bench and semi pilot scales are discussed in the final chapter This edition of the book has been organized in two parts In the first part a thorough discussion regarding reaction kinetics is presented In the second part basic equations are derived and used to represent the performances of batch and continuous ideal reactors isothermal and non isothermal reaction systems and homogeneous and heterogeneous reactor vessels as illustrated with several examples and exercises This textbook will be of great value to undergraduate and graduate students in chemical engineering as well as to graduate students in and researchers of kinetics and catalysis

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