

THE MOLE AND AVOGADRO'S NUMBER

Name _____

One mole of a substance contains Avogadro's Number (6.02×10^{23}) of molecules. (atoms)

How many molecules are in the quantities below?

1. 2.0 moles $x = 2.0 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{1.2 \times 10^{24} \text{ molecules}}$

2. 1.5 moles $x = 1.5 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{9.0 \times 10^{23} \text{ molecules}}$

3. 0.75 mole $x = 0.75 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{4.5 \times 10^{23} \text{ molecules}}$

4. 15 moles $x = 15 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{9.0 \times 10^{24} \text{ molecules}}$

5. 0.35 mole $x = 0.35 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{2.1 \times 10^{23} \text{ molecules}}$

How many moles are in the number of molecules below?

1. 6.02×10^{23} $x \text{ moles} = 6.02 \times 10^{23} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{1.0 \text{ mol}}$

2. 1.204×10^{24} $x \text{ moles} = 1.204 \times 10^{24} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{2.0 \text{ mol}}$

3. 1.5×10^{20} $x \text{ moles} = 1.5 \times 10^{20} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{2.5 \times 10^{-4} \text{ mol}}$ (0.00025 mol)

4. 3.4×10^{26} $x \text{ moles} = 3.4 \times 10^{26} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{565 \text{ mol}}$

5. 7.5×10^{19} $x \text{ moles} = 7.5 \times 10^{19} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{1.2 \times 10^{-4} \text{ mol}}$

Chemistry If8766 Answer And Work

Randal L. Henly



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Fundamentals of Chemistry Fred H. Redmore, 1979 Advanced Level Practical Work in Chemistry John Parkinson, 1991 *Practice Makes Perfect Chemistry Review and Workbook, Second Edition* Marian DeWane, Heather Hattori, 2018-12-28

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