

CHEM 102 – Recitation Week 7 – Answer Key

Ch. 16, 33. In each reaction, identify the Brønsted-Lowry acid, the Brønsted-Lowry base, the conjugate acid, and the conjugate base. Use symbols A, B, CA, and CB, respectively.



Concepts needed:

Brønsted-Lowry definition of acids and bases:

definition of conjugate acid-base pairs: their chemical formulas must differ by one proton, H^+ .

Note: proton = an atom of hydrogen that lost its electron, hence the H^+ symbol.

chemical equations

Answer:



A

B

CA

CB

Note: Brønsted-Lowry bases must have a lone pair of electrons available to bind the incoming proton by a covalent bond in which both shared electrons come from the base.

Recall from CHEM 101 how to determine the number of lone pairs in atoms of non-metals!

Such lone pairs are usually not shown in chemical formulas.



B

A

CA

CB



A

B

CA

CB



B

A

CA

CB

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M Tight



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