

Name Morgan WhiteSection 1B1

Post Lab Questions: Answer question on this sheet. Show all work.

1. 1.75 moles of H_2O_2 were placed in a 2.50 L reaction chamber at 307°C . After equilibrium was reached, 1.20 moles of H_2O_2 remained. Calculate the equilibrium constant, K_c , for the reaction.



2. Phosgene, COCl_2 , a poisonous gas, decomposes according to the equation



Calculate K_p for this reaction if $K_c = 0.083$ at 900°C .

$$K_p = K_c(\text{RT})^{\Delta n}$$

3. Calculate K_c for the reaction



given that the concentrations of each species at equilibrium are as follows: $[\text{HI}] = 0.85 \text{ mol/L}$, $[\text{I}_2] = 0.60 \text{ mol/L}$, $[\text{H}_2] = 0.27 \text{ mol/L}$.

4. Consider this reaction at equilibrium:



If the volume of the system is expanded at constant temperature, what change will occur in the position of the equilibrium?

Chemistry Post Lab Answers Exploring Equilibrium

R Bogdan



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Exploring General Chemistry in the Laboratory Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

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