

# ACIDS, BASES AND SALTS

|                                                                                                                                                |                                                                                                                                                              |
|------------------------------------------------------------------------------------------------------------------------------------------------|--------------------------------------------------------------------------------------------------------------------------------------------------------------|
| <b>Definition of an acid</b><br>An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions | <b>What causes the acidic properties of acids?</b><br>The hydrogen ions ( $H^+$ ) cause the acidic properties and these are formed in the presence of water. |
|------------------------------------------------------------------------------------------------------------------------------------------------|--------------------------------------------------------------------------------------------------------------------------------------------------------------|

|                                                                                    |                                                                                                                                                                |
|------------------------------------------------------------------------------------|----------------------------------------------------------------------------------------------------------------------------------------------------------------|
| <b>Another term to refer to an acid</b><br>An acid is called a <u>proton donor</u> | <b>Why an acid is also called a proton donor?</b><br>It's because an acid provides protons or hydrogen ions ( $H^+$ ) to other substances during the reaction. |
|------------------------------------------------------------------------------------|----------------------------------------------------------------------------------------------------------------------------------------------------------------|

|                                                            |                                                 |                                                                                                       |                                                                                                                 |
|------------------------------------------------------------|-------------------------------------------------|-------------------------------------------------------------------------------------------------------|-----------------------------------------------------------------------------------------------------------------|
| <b>Substance to which an acid provides protons</b><br>BASE | <b>A base is then called</b><br>PROTON ACCEPTOR | <b>Why the base is called a proton acceptor?</b><br>Because the base accepts hydrogen ions from acids | <b>Equation for the reaction between acid and base</b><br>$H^+_{(aq)} + OH^-_{(aq)} \longrightarrow H_2O_{(l)}$ |
|------------------------------------------------------------|-------------------------------------------------|-------------------------------------------------------------------------------------------------------|-----------------------------------------------------------------------------------------------------------------|

| Common laboratory acids |           | These three common laboratory acids are also called | Why these acids are also called mineral acids?                                                               |
|-------------------------|-----------|-----------------------------------------------------|--------------------------------------------------------------------------------------------------------------|
| Hydrochloric acid       | HCl       | MINERAL ACIDS                                       | They are derived from mineral salts i.e. chlorides for HCl, sulphates for $H_2SO_4$ and nitrates for $HNO_3$ |
| Sulphuric acid          | $H_2SO_4$ |                                                     |                                                                                                              |
| Nitric acid             | $HNO_3$   |                                                     |                                                                                                              |

| Other mineral acids known |           | Mineral salts from which the acid is derived | Organic acids known          | Naturally occurring acids               |
|---------------------------|-----------|----------------------------------------------|------------------------------|-----------------------------------------|
| Sulphurous acid           | $H_2SO_3$ | Derived from sulphites                       | Ethanoic acid ( $CH_3COOH$ ) | CITRIC ACID from lemons                 |
| Carbonic acid             | $H_2CO_3$ | Derived from carbonates                      |                              | TARTARIC ACID from grapes               |
| Phosphoric acid           | $H_3PO_4$ | Derived from phosphates                      | Methanoic acid ( $HCOOH$ )   | ACETIC ACID from vinegar                |
| Nitrous acid              | $HNO_2$   | Derived from nitrites                        |                              | LACTIC ACID from sour milk              |
|                           |           |                                              |                              | Hydrochloric acid from digestive juices |

|                                                                             |                                                                                                             |                                                                                                                                                                               |
|-----------------------------------------------------------------------------|-------------------------------------------------------------------------------------------------------------|-------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| <b>Whenever an acid is dissolved in water, it produces</b><br>HYDROGEN IONS | <b>Term given to the number of hydrogen ions produced by one molecule of an acid</b><br>BASICITY OF AN ACID | <b>Definition of basicity of an acid</b><br>BASICITY of an acid is the <u>number of hydrogen ions</u> produced by <u>one molecule</u> of an acid <u>in aqueous solution</u> . |
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# Chemistry Acid Bases And Salts Study Guide

**David Kirk**



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