General Chemistry II Chapter 17 Study Guide

Organic acids- carboxylic acids (R-COOH):

- Carboxylic acids (R-COOH): a weak organic acid; contains the carboxyl group (-COOH); never fully dissociates in water, an equilibrium process; ex. Formic acid, acetic acid, butyric acid, benzoic acid
 - R= rest of the molecule
- Strength of carboxylic acids: depends on the R group; more electronegative CI in R group shifts more electron density towards the R group and helps stabilize the structure after the H* ion is removed; thus, chloroacetic acid is the stronger acid in comparison to acetic acid (CH₂COOH)

Organic bases- amines:

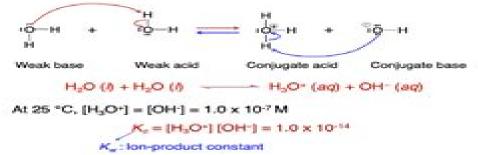
- Amines: organic bases; contains an N atom with a lone pair (ex. -NH, -NH₂); ex. Methyl amine, dimethyl amine, trimethyl amine; are weak bases:
 - Only N with lone pair can accept a proton
 - Equilibrium dissociation process
 - NH₄^{*} is an acid and can lose one H^{*} to be neutral

Organic acids and bases (things to keep in mind):

- Acid: H is the only acidic proton; ex. CH₃COOH becomes CH₃COO
- Amine: the proton H* is bonded to N via the lone pair of N; ex. CH₃NH₂ becomes CH₃NH₃*
- How about earbon?
 - Protons on C are not acidic; C cannot accept extra protons

Autoionization of water:

 Autoionization: in pure water, one water molecule can donate a proton to another water molecule



Chapter 17 Study Guide Chemistry Answer Key

Jiben Roy

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