General Chemistry II Chapter 13 Study Guide

- Raoult's Law: P = X + P =
 - P = vapor pressure of solvent "a" above a solution
 - P_a = vapor pressure of pure solvent "a"
 - X, = mole fraction of solvent "a" in the solution
 - X_a = moles of solvent "a" / moles of solution (solvent particles + solute particles)
- Volatile: substances that easily evaporate and exhibit a vapor pressure
- Nonvolatile: substances that have little tendency to evaporate
- Vapor pressure of the solution is the sum of the partial pressures of the solute and the solvent
 - If solute is non-volatile, vapor pressure of solution only comes from the solvent (P_{solution} < Pure _{solvent})
 - If solute is more volatile than solvent, Protettes > Pure notions

Review on boiling point:

- Boiling point: temperature where vapor pressure = atmospheric pressure
 - Vapor pressure of liquid increases as temperature increases
 - Starts boiling process when vapor pressure = atmospheric pressure
 - Solute in solution lowers the vapor pressure of solvent- vapor pressure lowering
 - A higher temperature needed for vapor pressure to hit atmospheric pressure
 - boiling point of solution is higher than boiling point of pure solvent-boiling point elevation

Boiling point elevation:

- boiling point elevation: the boiling point of a solution is always higher than that of the pure solvent, it depends on the type of solvent and the # of solute particles
- ΔT_b = iK_bm ; ΔT_b = T_b T_b
 - m= solute molality (moles solute / kg solvent)
 - K_b= molal boiling point elevation constant ('C / m); depends only on solvent
 - i= van't Hoff factor
 - The boiling point of solution
 - T_b = boiling point of pure solvent

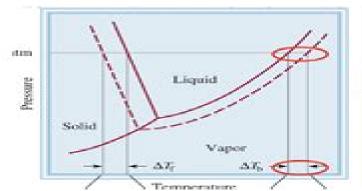


TABLE 13.2 Solvent	Depression Constants of Several Common Liquids			
	Normal Freezing Point (°C)*	K _e (°C/or)	Normal Boiling Point (°C)*	K _b (°C/m)
Water	0	1.86	100	0.52
Велисте	5.5	5.12	80.1	2.53
Ethanol	-117.3	1.99	78.4	1.22
Acetic neid	16.6	3.90	117.9	2.93
Cyclohexane	6.6	20.0	80.7	2.79

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