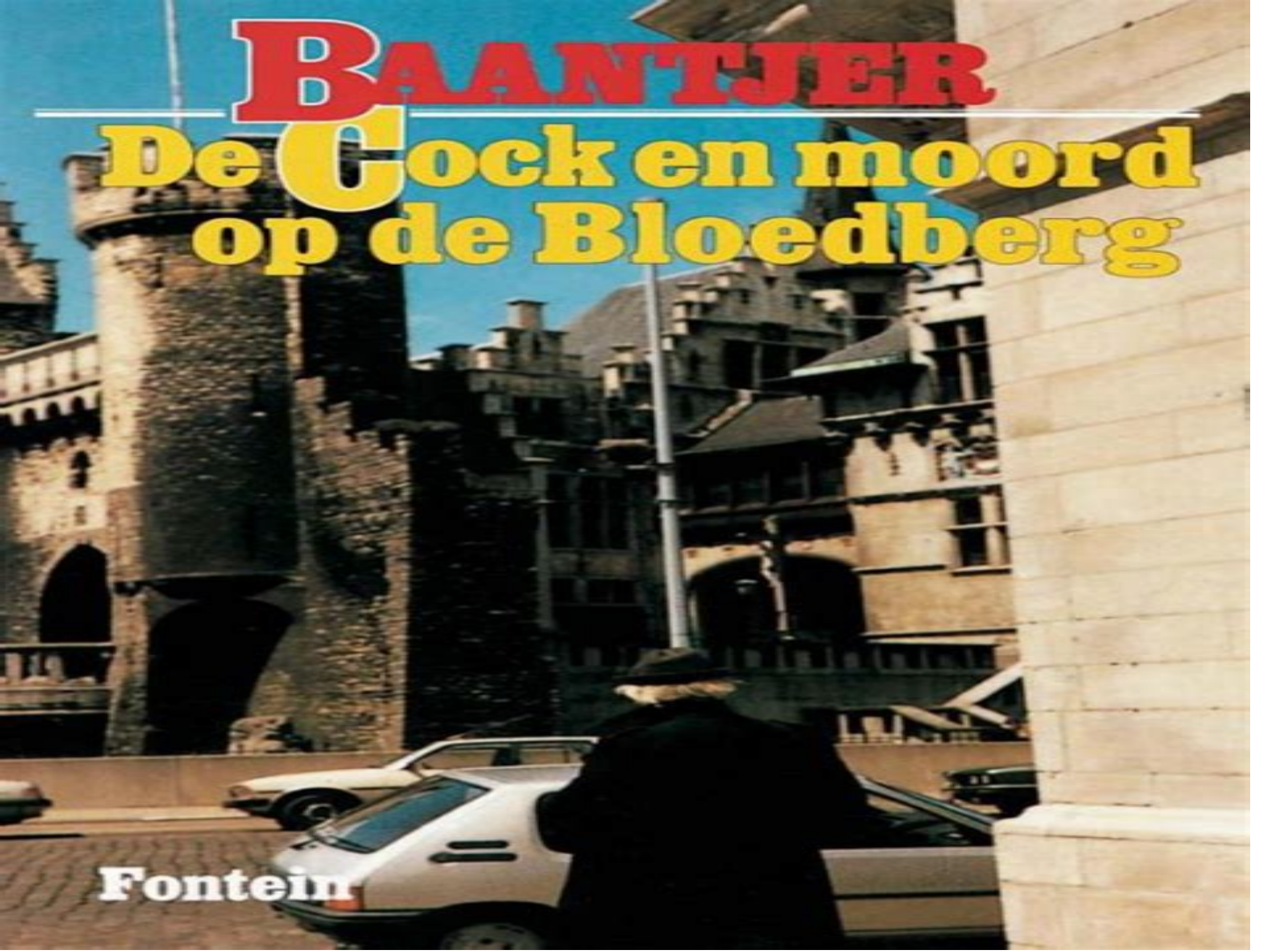




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web limiting reactant theoretical and percentage yields key 4 $\text{K}_2\text{S} + 2\text{H}_2\text{O} + 4\text{KOH} \rightarrow 3\text{O}_2 + \text{g}$ if a reaction vessel contains 0 mol K_2O and 0 mol H_2O what is the limiting reactant?

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 $\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $160.0 \text{ g } \text{Al} \times \frac{1 \text{ mol } \text{Al}}{27.0 \text{ g/mol}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol } \text{Al}} \times 101.96 \text{ g/mol } \text{Al}_2\text{O}_3 = 302.3 \text{ g } \text{Al}_2\text{O}_3$ c determine the percent yield $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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