

Question 8:

An atom X has 2, 8, 7 electrons in its shell. It combines with Y having 1 electron in its outermost shell.

(a) What type of bond will be formed between X and Y?

(b) Write the formula of the compound formed.

Solution 8:

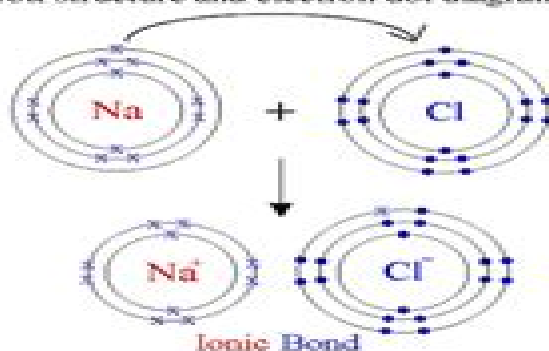
- (a) X has 7 electrons in its outermost shell and Y has only one electron in its outermost shell so Y loses its one electron and X gains that electron to form an ionic bond.
- (b) The formula of the compound would be XY.

Question 9:

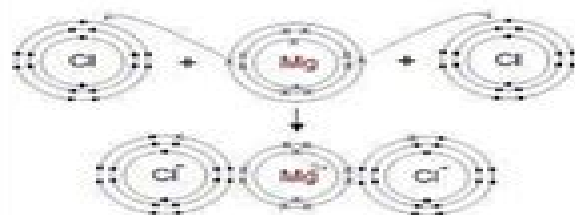
Draw orbit structure and electron dot diagram of NaCl, MgCl₂ and CaO.

Solution 9:

Orbit structure and electron dot diagram of NaCl:



Orbit structure and electron dot diagram of MgCl₂:



Orbit structure and electron dot diagram of CaO:

Concise Chemistry Chemical Bonding Chapter 2 Answers

CL Gary



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