

General Chemistry Placement Test, With Complete Solutions Updated | 2024\2024.

molecule two nonmetal atoms combined form a covalent compound called a
diatomic molecules H_2, O_2, F_2
Matter Anything that has mass and takes up space
Element A substance that cannot be broken down into simpler substances
Compound A pure substance made of two or more elements CHEMICALLY combined.
Compound Can be decomposed by chemically
Mixture A PHYSICAL blend of two or more components
(aq) aqueous mixture, a substance is dissolved in water
homogeneous mixture A mixture that is so evenly mixed that you can't see all of the different parts. Ex. Air, kool-aid
heterogeneous mixture A mixture in which different materials can be seen easily.
Ex. soil, chocolate chip cookies
filtration process to separate a solid "precipitate" from a liquid
sublimation phase change from (s) $\rightarrow$ (g)
empirical formula A formula showing the lowest whole number ratio of atoms in a compound. Can't be reduced.
solid A state of matter that has a definite shape and a definite volume
liquid A form of matter that flows, has constant volume, and takes the shape of its container
Bromine (Br) and Mercury (Hg) only two liquids on the periodic table
gas A state of matter with no definite shape or volume
gas will uniformly and completely fill a container
gases physical property most all nonmetals and group 18
physical property A characteristic of a substance that does not involve a chemical change.
Ex. melting/boiling point, solubility, density, color, or hardness.
properties used to identify or recognize matter (objects).
chemical property describes how matter reacts when it changes chemically into another substances. Ex. reactivity with acid, base or with air to rust, flammability, decomposition.
physical/phase change(s) $\rightarrow$ (l) $\rightarrow$ (g), diluting, crush, cut. Still same substance.
Density A physical property. Specific to every element.
Chemical symbol One capital letter, or a capital letter and a lower case letter together. Mass on upper left side, and atomic # on lower left side.
Conservation of Mass Equal # of atoms on both sides of a chemical equation
Democritus Greek philosopher that said all matter is made of tiny particles called "atomos" or atoms
JJ Thomson used the cathode ray tube to discover electrons

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