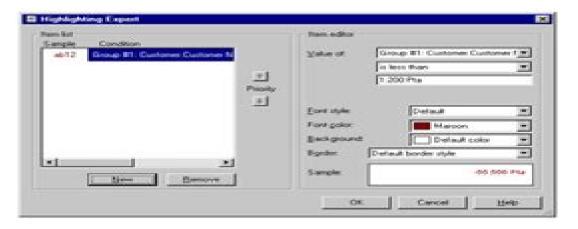


#### The Hightlighting Expert

Permite dar formato a los campos dependiendo de los valores que estos tengan. Es de gran utilidad cuando se presenta el caso donde en función de varias condiciones queremos dar formato a un campo, incluso estas condiciones tienen prioridades. La desventaja que posse, es que se debe configurar campo por campo, no podemos hacerlo al mismo tiempo para varios campos.

Por ejemplo: Si la deuda es mayor a 1000 mostrar en rojo, si la deuda es menor o igual a 1000 pero mayor a 100 mostrar en azul y si es menos o igual a 100 dejar como está. Para realizar este tipo de formato condicional, es una muy buena opción utilizar el Hightlighting Expert



#### Conditional Formatting Fórmulas

El formato condicional usa el editor de fórmulas, mediante el cual podemos crear fórmulas que determinen cuando y como queremos que aparezca el objeto.

Formato Absoluto es simplemente darle formato a los objetos con el editor de formato. Si se aplica formato condicional el formato absoluto es ignorado, siempre y cuando en la fórmula que se realice no utilice "DefaultAttribute", de lo contrario el atributo por defecto es el que se encuentra seteado.

Mediante esta opción si se pueden setear varios campos al mismo tiempo, incluso teniendo en cuenta el valor de cada campo seleccionado. Para poder hacer esto hay que utilizar la siguiente función: "CurrentFieldValue"

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by define shape and volume rigidity high density low compressibility the constituent particles atoms molecules or ions are closely packed and held together by strong interparticle forces types of solids

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chapter 1 the solid state talks about two main categories of solids which are crystalline solids and amorphous solids before understanding the differences between these solids let us learn their characteristic properties they have a definite shape volume and mass

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there are three states of matter solid liquid and gas in chapter 1 of ncert solutions for class 12 chemistry solid characteristics are studied solids are defined by their mass shape and volume intermolecular distance and force are both short yet intermolecular force is strong