

Chemistry Ch. 6 Study Guide

Bonds:

1. What characterizes a bond as Ionic?
2. What characterizes a bond as Covalent?

Ions

3. What is an ion?
4. How are Cations formed? What is the charge of a cation?
5. How are Anions formed? What is the typical charge of an anion?
6. Know how to use the periodic table to identify the type of ion formed and its charge.

Identifying a compound as ionic or molecular:

7. What type of elements are found in ionic compounds?
8. What type of elements are found in molecular compounds?
9. Ionic or Molecular?



Naming Compounds

10. How are Ionic Compounds named?
11. How are Molecular Compounds named?
12. Name the Following compounds:



Writing Formulas

13. Explain how to write a chemical formula for an ionic compound.
14. Explain how to write a formula for a molecular compound.
15. Identify the chemical formulas for the following compounds:

Sodium oxide

Carbon monoxide

Carbon tetrachloride

Magnesium nitride

***Be Familiar with the common metal ions with more than one charge (pg. 144) You do not have to memorize these, but you are responsible for knowing the charge on the ion when given a stock name.*

16. What is the charge of a lead (IV) ion?
17. What is the chemical formula for Lead (II) chloride?

Chapter 6 Study Guide 2

Kim Heldman



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