

Physics

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(Chapter 11) (Thermodynamics) (Exercises)

(Class 11)

Question 11.1:

A geyser heats water flowing at the rate of 3.0 litres per minute from 27 °C to 77 °C. If the geyser operates on a gas burner, what is the rate of consumption of the fuel if its heat of combustion is 4.0×10^4 J/g?

Answer 11.1:

Water is flowing at a rate of 3.0 litre/min.

The geyser heats the water, raising the temperature from 27°C to 77°C.

Initial temperature, $T_1 = 27^\circ\text{C}$ and Final temperature, $T_2 = 77^\circ\text{C}$

Rise in temperature, $\Delta T = T_2 - T_1 = 77 - 27 = 50^\circ\text{C}$

Heat of combustion = 4×10^4 J/g

Specific heat of water, $c = 4.2 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1}$

Mass of flowing water, $m = 3.0 \text{ litre/min} = 3000 \text{ g/min}$

Total heat used, $\Delta Q = mc \Delta T = 3000 \times 4.2 \times 50 = 6.3 \times 10^5 \text{ J/min}$

$\therefore$ Rate of consumption = $\frac{6.3 \times 10^5}{4 \times 10^4} = 15.75 \text{ g/min}$

Question 11.2:

What amount of heat must be supplied to 2.0×10^{-2} kg of nitrogen (at room temperature) to raise its temperature by 45 °C at constant pressure? (Molecular mass of $N_2 = 28$; $R = 8.3 \text{ J mol}^{-1} \text{ K}^{-1}$.)

Answer 11.2:

Mass of nitrogen, $m = 2.0 \times 10^{-2} \text{ kg} = 20 \text{ g}$ and Rise in temperature, $\Delta T = 45^\circ\text{C}$

Molecular mass of N_2 , $M = 28$

Universal gas constant, $R = 8.3 \text{ J mol}^{-1} \text{ K}^{-1}$

Number of moles, $n = \frac{m}{M} = \frac{2.0 \times 10^{-2} \times 10^3}{28} = 0.714$

Molar specific heat at constant pressure for nitrogen $C_p = \frac{7}{2}R = \frac{7}{2} \times 8.3 = 29.05 \text{ J mol}^{-1} \text{ K}^{-1}$

The total amount of heat to be supplied is given by the relation: $\Delta Q = nC_p \Delta T$

$= 0.714 \times 29.05 \times 45 = 933.38 \text{ J}$

Therefore, the amount of heat to be supplied is 933.38 J.

Question 11.3:

Explain why

- (a) Two bodies at different temperatures T_1 and T_2 if brought in thermal contact do not necessarily settle to the mean temperature $(T_1 + T_2)/2$.
- (b) The coolant in a chemical or a nuclear plant (i.e., the liquid used to prevent the different parts of a plant from getting too hot) should have high specific heat.
- (c) Air pressure in a car tyre increases during driving.
- (d) The climate of a harbour town is more temperate than that of a town in a desert at the same latitude.

Answer 11.3:

- (a) When two bodies at different temperatures T_1 and T_2 are brought in thermal contact, heat flows from the body at the higher temperature to the body at the lower temperature till equilibrium is achieved, i.e., the temperatures of both the bodies become equal. The equilibrium temperature is equal to the mean temperature $(T_1 + T_2)/2$ only when the thermal capacities of both the bodies are equal.

Chapter 24 Thermodynamics Exercises Answers

Mr. Ajeet Singh



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Physical Chemistry: Thermodynamics Horia Metiu, 2006-02-21 This is a new undergraduate textbook on physical chemistry by Horia Metiu published as four separate paperback volumes. These four volumes on physical chemistry combine a clear and thorough presentation of the theoretical and mathematical aspects of the subject with examples and applications drawn from current industrial and academic research. By using the computer to solve problems that include actual experimental data, the author is able to cover the subject matter at a practical level. The books closely integrate the theoretical chemistry being taught with industrial and laboratory practice. This approach enables the student to compare theoretical projections with experimental results, thereby providing a realistic grounding for future practicing chemists and engineers. Each volume of Physical Chemistry includes Mathematica and Mathcad Workbooks on CD ROM. Metiu's four separate volumes: Thermodynamics, Statistical Mechanics, Kinetics, and Quantum Mechanics, offer built-in flexibility by allowing the subject to be covered in any order. These textbooks can be used to teach physical chemistry without a computer, but the experience is enriched substantially for those students who do learn how to read and write Mathematica or Mathcad programs. A TI 89 scientific calculator can be used to solve most of the exercises and problems. *Student Solutions Manual to the Second Editions of Chemistry, Bailar ... [et Al.] and Chemistry with Inorganic Qualitative Analysis, Moeller ... [et Al.]*

John Williams, 1984 **Perspectives on Statistical Thermodynamics** Yoshitsugu Oono, 2017-12-14 This original text develops a deep conceptual understanding of thermal physics, highlighting the important links between thermodynamics and statistical physics, and examining how thermal physics fits within physics as a whole from an empirical perspective. The first part of the book is devoted to elementary mesoscopic topics such as Brownian motion, which leads to intuitive uses of large deviation theory, one of the pillars of modern probability theory. The book then introduces the key concepts behind statistical thermodynamics, and the final part describes more advanced and applied topics from thermal physics such as phase transitions and critical phenomena. This important subject is presented from a fresh perspective and in a highly pedagogical manner, with numerous worked examples and relevant cultural side notes throughout, making it ideal as either a textbook for advanced thermal physics courses or for self-study by undergraduate and graduate students in physics and engineering.

Physical Chemistry Peter Atkins, Julio de Paula, 2002 New edition of the overwhelmingly favorite text for the physical chemistry course. Solutions to Exercises Roxy Wilson, 2002-05 Full solutions to all end-of-chapter exercises in the text are provided. With an instructor's permission, this manual may be made available to students. Concepts of Physics Mr. Ajeet Singh, 2018-05-13 EduGorilla has its own publishing wing producing exam prep books, trade books, etc. *Solutions to Black Exercises* Roxy Wilson, 2005-05 Features detailed step-by-step solutions to the more than 1100 black-numbered end-of-character problems in Chemistry, the central science. **College Physics Textbook Equity Edition Volume 2 of 3: Chapters 13 - 24**

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