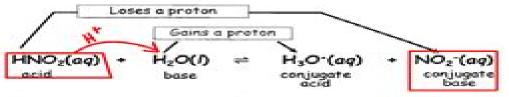
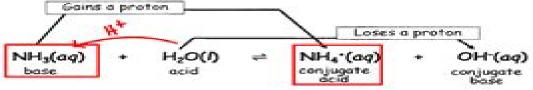
- B. Conjugate Acid-Base Pairs
- Every acid has a conjugate base and every base has a conjugate acid.
- A conjugate acid-base pair consists of two species that differ from each other by just one proton, H⁺.
- When a Brønsted acid donates a proton, what remains of the acid is known as a conjugate base.



- The two species HNO₂ and NO₂ are known as a conjugate acid-base pair or simply a conjugate pair.
- When a Brønsted base accepts a proton, the newly formed protonated species is known as a conjugate acid.



- The two species NH₃ and NH₄ are known as a conjugate acid-base pair or simply a conjugate pair.
 - PROBLEM-SOLVING CLASS ACTIVITY 1
 For the following reaction in aqueous solution, identify the conjugate acid-base pairs.

$$H_2PO_4^-(aq) + CO_3^2^-(aq) \implies HPO_4^2^-(aq) + HCO_3^-(aq)$$

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