



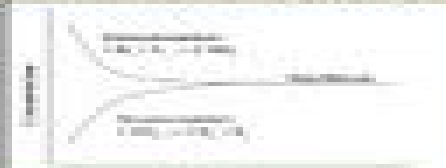
CHEMISTRY

A CHEAT SHEET

EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

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What is Equilibrium? Equilibrium is a state of balance between two opposing forces or processes. In chemistry, it refers to a reaction where the forward and reverse rates are equal, resulting in no net change in the concentrations of reactants and products over time. At equilibrium, the system is dynamic, meaning that the reaction is still occurring, but the concentrations of reactants and products remain constant. Equilibrium Constant (K) For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant (K) is defined as: $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ where [A], [B], [C], and [D] are the molar concentrations of the reactants and products at equilibrium.	Equilibrium Constant (K) For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant (K) is defined as: $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ where [A], [B], [C], and [D] are the molar concentrations of the reactants and products at equilibrium.	 The graph shows two curves: one for reactants (decreasing) and one for products (increasing). They meet at a point where the concentrations are constant, indicating equilibrium.								
Equilibrium Constant (K_c) For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant (K_c) is defined as: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ where [A], [B], [C], and [D] are the molar concentrations of the reactants and products at equilibrium.	<table><tr><th>Equilibrium Constant (K_c)</th><th>Meaning</th></tr><tr><td>$K_c > 1$</td><td>Products are favored at equilibrium.</td></tr><tr><td>$K_c < 1$</td><td>Reactants are favored at equilibrium.</td></tr><tr><td>$K_c = 1$</td><td>Reactants and products are present in equal concentrations at equilibrium.</td></tr></table>	Equilibrium Constant (K _c)	Meaning	$K_c > 1$	Products are favored at equilibrium.	$K_c < 1$	Reactants are favored at equilibrium.	$K_c = 1$	Reactants and products are present in equal concentrations at equilibrium.	
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Equilibrium Constant (K_p) For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant (K_p) is defined as: $K_p = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$ where P_A, P_B, P_C, and P_D are the partial pressures of the reactants and products at equilibrium.	<table><tr><th>Equilibrium Constant (K_p)</th><th>Meaning</th></tr><tr><td>$K_p > 1$</td><td>Products are favored at equilibrium.</td></tr><tr><td>$K_p < 1$</td><td>Reactants are favored at equilibrium.</td></tr><tr><td>$K_p = 1$</td><td>Reactants and products are present in equal concentrations at equilibrium.</td></tr></table>	Equilibrium Constant (K _p)	Meaning	$K_p > 1$	Products are favored at equilibrium.	$K_p < 1$	Reactants are favored at equilibrium.	$K_p = 1$	Reactants and products are present in equal concentrations at equilibrium.	
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Le Chatelier's Principle When a system at equilibrium is subjected to a change in concentration, pressure, or temperature, the system will adjust itself to counteract the change and restore a new equilibrium. For example, if the concentration of a reactant is increased, the system will shift to the right, favoring the products, to reduce the concentration of the reactant.	<table><tr><th>Equilibrium Constant (K_p)</th><th>Meaning</th></tr><tr><td>$K_p > 1$</td><td>Products are favored at equilibrium.</td></tr><tr><td>$K_p < 1$</td><td>Reactants are favored at equilibrium.</td></tr><tr><td>$K_p = 1$</td><td>Reactants and products are present in equal concentrations at equilibrium.</td></tr></table>	Equilibrium Constant (K _p)	Meaning	$K_p > 1$	Products are favored at equilibrium.	$K_p < 1$	Reactants are favored at equilibrium.	$K_p = 1$	Reactants and products are present in equal concentrations at equilibrium.	
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Le Chatelier's Principle: The effect of changing conditions on equilibrium.										
Concentration	- Increasing the concentration of a reactant will shift the equilibrium to the right, favoring the products. - Increasing the concentration of a product will shift the equilibrium to the left, favoring the reactants.	For a general reaction: $aA + bB \rightleftharpoons cC + dD$ Increasing [A] or [B] will increase [C] and [D].								
Pressure	- Increasing the pressure will shift the equilibrium to the side with the fewest moles of gas. - Decreasing the pressure will shift the equilibrium to the side with the most moles of gas.	For a general reaction: $aA + bB \rightleftharpoons cC + dD$ If $a + b > c + d$, increasing pressure will favor the products.								
Temperature	- Increasing the temperature will shift the equilibrium to the side that is endothermic (absorbs heat). - Decreasing the temperature will shift the equilibrium to the side that is exothermic (releases heat).	For a general reaction: $aA + bB \rightleftharpoons cC + dD$ If the reaction is endothermic, increasing temperature will favor the products.								

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