Name KEY Test Date Period

CP Chemistry - Final Exam Review

Show your work and round using significant figure rules for all the problems.

- What is chemistry?
 - Chemistry is the study of matter and the changes it undergoes.
- Write the following in scientific notation:
 - a. .005 5 x 10⁻³

c. 53099000 5.3099 x 10⁷

b. 5050 5,05 x 10³

d. 0.00053030 5.3030 x 10⁻⁴

- Write the following in standard notation:
 - a. 1.5x10³ 1500

c. 4x10⁻⁵ 0.00004

b. 3.75 ×10⁻² 0.0375

- d. 9.12 x10⁷ 91200000
- How many significant figures are in the following:
 - a. 0.02
- 1

- c. 501
- -9

b. 501.0

- d. 0.020
- . 2

- Calculate and report in correct significant figures:
 - a. 52.568 cm + 689.45 cm + 55.5694 cm = 797.59 cm
 - b. 8426, 200 g ÷ 254 mol = 33.2 g /mol
- You measure a Calcium cube to have 2.01 cm sides and a mass of 11.6 grams:
 - a. What is the density as you calculate it?

$$y = s^3 = (2.01 \text{ cm})^3 = 8.12 \text{ cm}^3$$

m 11.6 g

$$d = \frac{m}{v} = \frac{11.6 \text{ g}}{8.12 \text{ cm}^3} = 1.43 \text{ g/cm}^3$$

b. If calcium's accepted density is 1.54 g/cm³, what is your percent error?
11.54 g/cm³ = 1.43 g/cm³

$$\%_{error} = \frac{|1.54 \text{ g/cm}^3 - 1.43 \text{ g/cm}^3|}{1.54 \text{ g/cm}^3} \times 100 = 7.14 \%$$

- Perform the following conversions:
 - 53.1 mm to m

53.1 mm x
$$\frac{1 \text{ m}}{10^3 \text{ mm}} = 0.0531 \text{ m}$$

b. 18.4 kg to dag

$$18.4 \text{ kg x} \frac{1 \text{ x } 10^{-1} \text{ dag}}{1 \text{ x } 10^{-3} \text{ kg}} = 1840 \text{ dag}$$

c. 93.6 mi to km (1 mi = 1.6 km)

93.6 mi x
$$\frac{1.6 \text{ km}}{1 \text{ mi}}$$
 = 150, km

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JE Gale

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- 1. Understanding the eBook Chemistry Semester Exam Study Guide Answers
 - The Rise of Digital Reading Chemistry Semester Exam Study Guide Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Chemistry Semester Exam Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Semester Exam Study Guide Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chemistry Semester Exam Study Guide Answers
 - Personalized Recommendations
 - Chemistry Semester Exam Study Guide Answers User Reviews and Ratings
 - Chemistry Semester Exam Study Guide Answers and Bestseller Lists
- 5. Accessing Chemistry Semester Exam Study Guide Answers Free and Paid eBooks
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 - Highlighting and Note-Taking Chemistry Semester Exam Study Guide Answers
 - Interactive Elements Chemistry Semester Exam Study Guide Answers
- 8. Staying Engaged with Chemistry Semester Exam Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Semester Exam Study Guide Answers
- 9. Balancing eBooks and Physical Books Chemistry Semester Exam Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Semester Exam Study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Semester Exam Study Guide Answers
 - Setting Reading Goals Chemistry Semester Exam Study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Semester Exam Study Guide Answers
 - Fact-Checking eBook Content of Chemistry Semester Exam Study Guide Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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