

How to Pass the Chemistry Regents

84	Periodic Table	Periodic Table Arrangement	increasing atomic number	8.21
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5. The elements on the Periodic Table of the Elements are arranged in order of increasing:

- (1) atomic mass (3) atomic number
(2) formula mass (4) oxidation number

4. The elements on the Periodic Table of the Elements are arranged in order of increasing:

- (1) atomic mass (3) mass number
(2) atomic number (4) oxidation state

6. The elements on the Periodic Table of the Elements are arranged in order of increasing:

- (1) atomic number
(2) mass number
(3) number of neutrons
(4) number of valence electrons

85	Periodic Table	Properties of metals/nonmetals	good conductors of electricity, malleable (deformable) poor conductors of electricity, brittle (breaks easily)	8.22
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3. Which properties are characteristic of Group 2 elements at STP?

- (1) good electrical conductivity and electronegativity less than 1.7
(2) good electrical conductivity and electronegativity greater than 1.7
(3) poor electrical conductivity and electronegativity less than 1.7
(4) poor electrical conductivity and electronegativity greater than 1.7

7. Which element is malleable at STP?

- (1) chlorine (3) helium
(2) copper (4) sulfur

Metals and nonmetals have different chemical and physical properties. Some of these properties are shown in the table below.

Some Physical and Chemical Properties of Rubidium and Sulfur

Rubidium	Sulfur
shiny silver metal	crystalline, yellow solid
reacts with water to form hydrogen gas and rubidium hydroxide	combusts in air with a blue flame
melts with copper at low pH	melts at room temperature
specific heat is 0.03 J/g°C	specific heat is 0.21 J/g°C

13. A solid sample of copper is an excellent conductor of electric current. Which type of chemical bonds are in the sample?

- (1) ionic bonds
(2) metallic bonds
(3) trigonal covalent bonds
(4) polar covalent bonds

[10] Compare the electrical conductivity of these two elements at STP. (1)

10. The solid below has properties of selected elements at room temperature:

Properties of Selected Elements at Room Temperature

Element	Density (g/cm ³)	Malleability	Conductivity
metal	2.00	yes	good
gas	0.001	yes	good
solid	2.00	no	poor
liquid	1.00	yes	good

Based on the table, which statement describes how two of these elements can be differentiated from each other?

- (1) Compare the different mass numbers of these two elements.
(2) Compare the different number of neutrons in the nuclei of these two elements.
(3) Rubidium can be differentiated from hydrogen based on conductivity.
(4) Sulfur can be differentiated from sulfur based on malleability.

Chemistry Regents Review April 2008

Clifford Lane



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