#### Chemistry diagnostic test (Topic 12)

B – the p<sup>4</sup> config. of S is less stable, electron-electron repulsion.

## tlt t

- A B is lower than Be because two s electrons shield the 2p electrons
- E Au and Ag both have full d orbitals(10electrons) (s<sup>1</sup>)
- E electronegativity increases from lower left to upper right (F-most electronegative)
- B sulfur's oxidation numbers: 2-, 2+, 4+, 6+
- A atomic radii increase right to left and top to bottom
- C cations are smaller, the higher the charge the smaller the cation
- A Group 1 has low 1<sup>st</sup> ionization energy, but 2<sup>nd</sup> is pulling an inner core electron(big jump)
- E equal moles 0.25
- D greatest molar mass
- 11. D greatest mass/volume
- 12. E all same temp
- 13. E Phosphorus high reactivity with oxygen not found free in nature
- 14. B not all have a 2+ oxidation
- A all substances contain London dispersion forces
- 16. D hydrogen must be bonded to nitrogen, oxygen or fluorine
- 17. A propane (molar mass 44) dispersion forces only lowest temp Ethanol (molar mass – 46) also has hydrogen bonding Water (molar mass – 18) very strong hydrogen bonds
- 18. C
- A balanced equation is
  - $Na2CO3 + 2HC1 \rightarrow 2NaC1 + CO2 + H20$
  - Limiting is sodium carbonate
- 20. C = (0.300 moles/L)(.300 L) = 0.0900 mol CI- ion requires 1/3 as much aluminum chloride or 0.0300 mol
  - 0.0300 mol/0.500Mormol/L = 0.0600L = 60.0 ml
- 21. B 0.500 M x 0.250L x 200g/mol = 25.0 g
- 22. D- 2 C3H8O + 9O2 → 6CO2 + 8H2O
- C 3.00 carbon = ½ mole of C therefore ½ mol of Z combines 0.5 mol Z x 36.0 g/mol(3timesC) = 18.0 g Z
- 24. E 4 moles of O
- C an alkali metal reacts with water to produce hydrogen gas and an aqueous alkali metal hydroxide
- 26. D H2SO4 + NaOH → NaHSO4 + H2O (diprotic acid and monobasic base
- 27. E-X has a 3+ charge
- 28. E
- B 1 mole of magnesium requires 2 moles of HCl therefore HCl is limiting
- 30. D a formation reaction produces 1 mole of a product

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#### Table of Contents Chemistry Answer Key Diagnostic Test Topic 4

- 1. Understanding the eBook Chemistry Answer Key Diagnostic Test Topic 4
  - The Rise of Digital Reading Chemistry Answer Key Diagnostic Test Topic 4
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Chemistry Answer Key Diagnostic Test Topic 4
  - Exploring Different Genres
  - $\circ\,$  Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Answer Key Diagnostic Test Topic 4
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chemistry Answer Key Diagnostic Test Topic 4

- Personalized Recommendations
- Chemistry Answer Key Diagnostic Test Topic 4 User Reviews and Ratings
- Chemistry Answer Key Diagnostic Test Topic 4 and Bestseller Lists
- 5. Accessing Chemistry Answer Key Diagnostic Test Topic 4 Free and Paid eBooks
  - Chemistry Answer Key Diagnostic Test Topic 4 Public Domain eBooks
  - Chemistry Answer Key Diagnostic Test Topic 4 eBook Subscription Services
  - Chemistry Answer Key Diagnostic Test Topic 4 Budget-Friendly Options
- 6. Navigating Chemistry Answer Key Diagnostic Test Topic 4 eBook Formats
  - o ePub, PDF, MOBI, and More
  - Chemistry Answer Key Diagnostic Test Topic 4 Compatibility with Devices
  - Chemistry Answer Key Diagnostic Test Topic 4 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Answer Key Diagnostic Test Topic 4
  - Highlighting and Note-Taking Chemistry Answer Key Diagnostic Test Topic 4
  - Interactive Elements Chemistry Answer Key Diagnostic Test Topic 4
- 8. Staying Engaged with Chemistry Answer Key Diagnostic Test Topic 4
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Answer Key Diagnostic Test Topic 4
- 9. Balancing eBooks and Physical Books Chemistry Answer Key Diagnostic Test Topic 4
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Answer Key Diagnostic Test Topic 4
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Answer Key Diagnostic Test Topic 4
  - Setting Reading Goals Chemistry Answer Key Diagnostic Test Topic 4
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Answer Key Diagnostic Test Topic 4

- Fact-Checking eBook Content of Chemistry Answer Key Diagnostic Test Topic 4
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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