

- 4 Hydrogen and chlorine are reacted together to form hydrogen chloride.



- (a) Calculate the bond enthalpy for the H–Cl bond using the information in the table below.

| bond  | bond enthalpy<br>/ $\text{kJ mol}^{-1}$ |
|-------|-----------------------------------------|
| H–H   | +436                                    |
| Cl–Cl | +243                                    |

Bonds broken

436

- 243

= 679

- 679

Bonds made 2(HCl)

Bonds Broken + Bonds Made = -184

-679 + 2HCl = -184

2HCl = 495

HCl = 247.5

bond enthalpy = .....  $\text{kJ mol}^{-1}$  [2]

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