

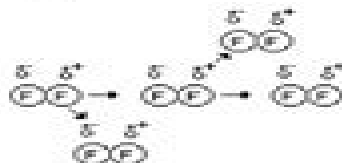
CHAPTER 11 PRACTICE QUESTIONS KEY

INSTRUCTORS: Do not release this key until noon 4/22/18.

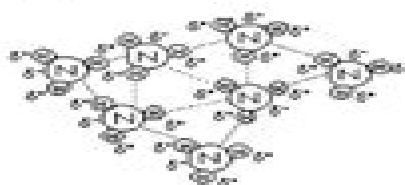
Use the following to answer questions 1-3:

Consider the representations below to answer the next three questions.

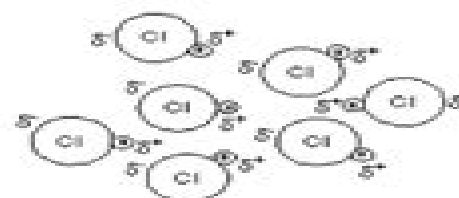
(I)



(II)



(III)



(III)

- How many of the following statements are correct concerning drawing I?
 - Each molecule induces a dipole onto the next molecule in close proximity.
 - The phenomenon shown is relatively weak and short-lived.
 - C_6H_{18} contains this type of interaction.
 - The forces that exist in this example are London dispersion forces.

a) 0 b) 1 c) 2 d) 3 e) 4
- Which drawing *best* represents the interactions in a sample of HF?

a) I
b) **II**
c) III
d) I, II
e) all of the above
- Which of the following statements are incorrect concerning drawing III?

a) Electrostatic interactions exist between the molecules.
b) The molecules find the best compromise between attraction and repulsion.
c) **These molecules exhibit ionic bonding.**
d) OCS exhibits this type of interaction.
e) Two of the above statements are incorrect.
- Which of the following is the correct order of boiling points for KNO_3 , CH_3OH , C_2H_6 , Ne?

a) $Ne < CH_3OH < C_2H_6 < KNO_3$
b) $KNO_3 < CH_3OH < C_2H_6 < Ne$
c) $Ne < C_2H_6 < KNO_3 < CH_3OH$
d) **$Ne < C_2H_6 < CH_3OH < KNO_3$**
e) $C_2H_6 < Ne < CH_3OH < KNO_3$
- What is the strongest intermolecular force (IMF) found in a sample of thousands of HBr molecules?

a. London b. **dipole dipole** c. nonpolar d. H bond force e. polar

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Sebastian Brünink



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