Chapter 1 - Introduction

Classifications of Matter – 1.3

- Matter: anything that occupies space and has mass
- Chemistry: the study of matter and its change
- Atom: smallest distinctive unit
- Molecule: two or more atoms together
- Substance: group of identical atoms or molecules.
 - Element: atoms are identical
 - Compound: contains different atoms
- Homogenous Mixture: uniform, seems as one thing
- Heterogeneous Mixture: non-uniform, able to see the elements

ments Pure Substance Mixtures

Physical and Chemical Properties of Matter - 1.4

- Physical Property: measurable, describes system stateCompounds Elements Heterogeneous Homogeneous

 Mixture Mixture
 - "Air is a gas," "boiling point of water is 100 degrees," "Sulfur crystals are yellow"
- Chemical Property: chemical reaction, compound cannot be restored without a change
 - Burning, explosions, decomposition, rusting
 - i.e. "Burning Hydrogen in Oxygen to form water," "milk turning sour"

Measurement – 1.5

- SI base units of Measurement:
 - Length meter (m)
 - Mass kilogram (kg)
 - Time second (s)
 - Temperature Kelvin (K)
 - Amount of a substance mole (mol)
- Temperature Scales:
 - Kelvin: absolute temperature scale, 273° C = 1 K

Handling Number - 1.6

- Accuracy: Correctness How close the results were to where you intended them to be
- Precision: Exactness How close the results were in relation to other results / trials
- Scientific Notation: Expresses a number as a product of a number between 1 and 10 and the appropriate power of 10
 - All non-zeros are significant (1-9)
 - A zero between non-zeros is significant (209)
 - Final zeros to the right of the decimal are significant (10.0)
 - Initial zeros are not significant (0.47)
 - Final zeros in a number with no decimal may or may not be significant (9,500)
- Adding and subtracting: same number of digits past the decimal point as the number used with the least precision – the least amount of significant figures
 - If greater than 5, round up
- Multiplication and Division: least number or significant figures
- Exact Numbers: have no effect on significant figures i.e. 60 seconds in a minute
- Volume (V): Length cubed, mL = 1 cm³

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