

1 Atoms, molecules and stoichiometry

DEFINITION: Relative atomic mass (A_r) is the **average mass** of one atom compared to one twelfth of the mass of one atom of carbon-12

DEFINITION: Relative isotopic mass is the **mass** of one isotope compared to one twelfth of the mass of one atom of carbon-12

DEFINITION: Relative molecular mass (M_r) is the **weighted average mass** of a molecule compared to one twelfth of the mass of one atom of carbon-12

DEFINITION: Relative formula mass (M_r) is the **weighted average masses** of the **formula units** compared to one twelfth of the mass of one atom of carbon-12

The term molecule should only be used for covalent molecules. For ionic substances use the term relative formula mass. They are, however, calculated in the same way.

Avogadro's Number

There are 6.022×10^{23} atoms in 12 grams of carbon-12. Therefore explained in simpler terms 'One mole of any specified entity contains 6.022×10^{23} of that entity'.

The mole

The mole is the key concept for chemical calculations

DEFINITION: The mole is the amount of substance in grams that has the same number of particles as there are atoms in 12 grams of carbon-12.

For pure solids, liquids and gases

$$\text{moles} = \frac{\text{mass}}{M_r}$$

Unit of Mass: grams

Unit of moles : mol

Example 1: What is the number of moles in 35.0g of CuSO_4 ?

$$\begin{aligned}\text{moles} &= \text{mass}/M_r \\ &= 35.0 / (63.5 + 32.0 + 16.0 \times 4) \\ &= 0.219 \text{ mol}\end{aligned}$$

Relative Formula mass (M_r) for a compound can be calculated by adding up the relative atomic masses (from the periodic table) of each element in the compound
eg $\text{CaCO}_3 = 40.1 + 12.0 + 16.0 \times 3 = 100.1$

Many questions will involve changes of units

1000 mg = 1g

1000 g = 1kg

1000kg = 1 tonne

Example 2: What is the number of moles in 75.0mg of $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$?

$$\begin{aligned}\text{moles} &= \text{mass}/M_r \\ &= 0.075 / (40 + 32.0 + 16.0 \times 4 + 18.0 \times 2) \\ &= 4.36 \times 10^{-4} \text{ mol}\end{aligned}$$

Avogadro's Constant can be used for atoms, molecules and ions

1 mole of copper atoms will contain 6.022×10^{23} atoms

1 mole of carbon dioxide molecules will contain 6.022×10^{23} molecules

1 mole of sodium ions will contain 6.022×10^{23} ions

$$\text{No of particles} = \text{moles of substance (in mol)} \times \text{Avogadro's constant (L)}$$

Example 3 : How many atoms of Tin are there in a 6.00 g sample of Tin metal?

$$\begin{aligned}\text{moles} &= \text{mass}/A_r \\ &= 6.00 / 118.7 \\ &= 0.05055 \text{ mol} \\ \text{Number atoms} &= \text{moles} \times 6.022 \times 10^{23} \\ &= 0.05055 \times 6.022 \times 10^{23} \\ &= 3.04 \times 10^{22}\end{aligned}$$

Example 4 : How many chloride ions are there in a 25.0 cm³ of a solution of magnesium chloride of concentration 0.400 mol dm⁻³?

$\text{moles} = \text{concentration} \times \text{Volume}$

$$\text{MgCl}_2 = 0.400 \times 0.0250$$

$$= 0.0100 \text{ mol}$$

$$\begin{aligned}\text{moles of chloride ions} &= 0.0100 \times 2 \\ &= 0.0200\end{aligned}$$

$$\text{Number ions of Cl}^- = \text{moles} \times 6.022 \times 10^{23}$$

$$= 0.0200 \times 6.022 \times 10^{23}$$

$$= 1.20 \times 10^{22} \text{ (to 3 sig fig)}$$

There are two moles of chloride ions for every one mole of MgCl_2

Chemistry Study Guide Answers Atoms And Molecules

Arshad Iqbal



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