

- b. the relative number of atoms in a molecule
 c. one formula mass of water
 d. one gram of water
6. A gas has a density of 0.717 g/l at STP. What is the mass of one mole of this gas?
 a. 8.0 g
 b. 16.1 g
 c. 28.1 g
 d. 4.0 g
7. The molecular mass of CO_2 is 44; therefore, at STP
 a. 1 liter of CO_2 weighs 44 g
 b. 11.2 liters of CO_2 weigh 22 g
 c. the molar volume of CO_2 is 44
 d. 22.4 g. of CO_2 have a volume of 44 liters
8. The formula, which gives the simple ratio of each kind of atoms present in the molecule of compound, is called _____
 a. Molecular Formula
 b. Empirical Formula
 c. Structural Formula
 d. formula unit
9. Smallest particle of an element which may or may not have independent existence is known as.
 a. A molecule
 b. An ion
 c. An atom
 d. An electron
10. Atoms and molecules can either gain or lose electrons, forming charged particles called:
 a. Positrons
 b. Atoms
 c. Electrons
 d. Ions
11. The ratio of one mole of an element compared with the mass of one mole of another element is called
 a. Atomic weight ratio
 b. Atomic mass
 c. Molecular mass
 d. Relative atomic mass
12. How many moles are present in one mole of water?
 a. 6.023×10^{23}
 b. 1
 c. 6.023×10^{24}
 d. 2
13. The branch of chemistry which deals with quantitative relationship between reactants and products in a balanced chemical equation is called
 a. Analytical chemistry
 b. Organic chemistry
 c. Biochemistry
 d. Stoichiometry
14. A balanced chemical reaction
 a. Which provides quantitative relation of number of product
 b. Which provides quantitative relation of number of product
 c. Does not affect the amount of product
 d. Which is present in excess
15. The amount of product obtained in a balanced reaction is called
 a. Yield of a reaction
 b. Product of a reaction
 c. Consumption of a reaction
 d. Mass of atoms

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