

## Chemistry 338 2<sup>nd</sup> Semester Final Exam Review Sheet

### Unit 1 - Matter & Measurement – Chapters 1 & 2

1. A student experimentally determines the density of an object to be 4.2 g/mL. If the accepted value is 5.0 g/mL, calculate the % error.

$$\% \text{ error} = \frac{\text{experimental} - \text{accepted}}{\text{accepted}} \times 100$$

$$\% \text{ error} = \frac{4.2 - 5.0}{5.0} \times 100 = -16\% \text{ error}$$

### Unit 2 – Periodic Table – Chapter 5

2.

| Element      | Symbol                                       | Atomic # | Protons | Neutrons | Electrons | Mass # | Charge |
|--------------|----------------------------------------------|----------|---------|----------|-----------|--------|--------|
| Tungsten ion | <sup>184</sup> <sub>74</sub> W <sup>+2</sup> | 74       | 74      | 110      | 72        | 184    | +2     |

3. Using the following atomic radius values, predict the atomic radius of potassium (K):  
 Li – 145 pm, Na – 180 pm, K - ?, Rb – 235 pm  
**K – 220 pm**

### Unit 3 – Reactions/Compounds – Chapters 7 & 8

4. Classify the compound as ionic or covalent. Name the compound
- Fe<sub>2</sub>O<sub>3</sub> - iron (III) oxide (ionic - roman numerals for transition metals, change ending to -ide)
  - Cl<sub>2</sub>O<sub>8</sub> - dichlorine octaoxide (covalent - use prefixes, change ending to -ide)
5. Classify the compound as ionic or covalent. Write the formula
- Trisulfur Dinitride – S<sub>3</sub>N<sub>2</sub> (covalent)
  - Zinc Hydroxide – Zn(OH)<sub>2</sub> (ionic – look up ion charges and criss-cross) Zn<sup>+2</sup> OH<sup>-1</sup>
  - Chromium (III) nitrate – Cr(NO<sub>3</sub>)<sub>3</sub> (ionic – criss-cross) Cr<sup>+3</sup> NO<sub>3</sub><sup>-1</sup>
6. Write the equation, balance it, and classify the type of reaction:  
 Iron (II) sulfate reacts with sodium hydroxide to yield iron(II) hydroxide and sodium sulfate
- $$1 \text{ FeSO}_4 + 2 \text{ NaOH} \rightarrow 1 \text{ Fe(OH)}_2 + 1 \text{ Na}_2\text{SO}_4$$
- double displacement (two ionic compounds reacting)
7. Balance the equation and classify the type of reaction:
- $$1 \text{ Cu} + 2 \text{ AgNO}_3 \rightarrow 1 \text{ Cu(NO}_3)_2 + 2 \text{ Ag}$$
- Single displacement (element and a compound reacting)

### Unit 4 – Moles – Chapter 3

8. Calculate the molar mass of Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>.

$$\begin{aligned} \text{Fe (2)(55.9 g/mol)} &= 111.8 \text{ g/mol} \\ \text{S (3)(32.1 g/mol)} &= 96.3 \text{ g/mol} \\ + \text{O (12)(16.0 g/mol)} &= +192 \text{ g/mol} \\ &= 400.1 \text{ g/mol} \end{aligned}$$

# Chemistry Final Exam Review Sheet

**JR Anderson**



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**Senior Year**  
Judith P. Foard,2013-01-02 Ruth Ainsworth is a happy confident teen living in Massachusetts in 1947 She has a strong circle of friends and a secure family and she excels at her studies She is full of optimism for her senior year mostly because shes dating Rex Gardner the star of the Hampden High School football team They start their last school year and everything is going well for Ruthuntil Rex suddenly breaks up with her and begins dating a junior cheerleader Ruth is heartbroken but she

fight off the pain by focusing on school and the girls basketball team When Ruth meets Maurice Langlois a French speaking Canadian immigrant he helps her forget about Rex As their love grows however they realize they will soon be separated by hundreds of miles when they leave for college Their attempt to resolve problems related to their separation gets even more complicated when Rex makes a surprising reappearance in Ruth's life Will Ruth and Maurice's love for each other survive

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## **Table of Contents Chemistry Final Exam Review Sheet**

1. Understanding the eBook Chemistry Final Exam Review Sheet
  - The Rise of Digital Reading Chemistry Final Exam Review Sheet
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Final Exam Review Sheet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Final Exam Review Sheet
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Final Exam Review Sheet
  - Personalized Recommendations
  - Chemistry Final Exam Review Sheet User Reviews and Ratings
  - Chemistry Final Exam Review Sheet and Bestseller Lists

5. Accessing Chemistry Final Exam Review Sheet Free and Paid eBooks
  - Chemistry Final Exam Review Sheet Public Domain eBooks
  - Chemistry Final Exam Review Sheet eBook Subscription Services
  - Chemistry Final Exam Review Sheet Budget-Friendly Options
6. Navigating Chemistry Final Exam Review Sheet eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Final Exam Review Sheet Compatibility with Devices
  - Chemistry Final Exam Review Sheet Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Final Exam Review Sheet
  - Highlighting and Note-Taking Chemistry Final Exam Review Sheet
  - Interactive Elements Chemistry Final Exam Review Sheet
8. Staying Engaged with Chemistry Final Exam Review Sheet
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Final Exam Review Sheet
9. Balancing eBooks and Physical Books Chemistry Final Exam Review Sheet
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Final Exam Review Sheet
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Final Exam Review Sheet
  - Setting Reading Goals Chemistry Final Exam Review Sheet
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Final Exam Review Sheet
  - Fact-Checking eBook Content of Chemistry Final Exam Review Sheet
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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