

CHEMISTRY EQUATIONS & ANSWERS

HOW TO BALANCE CHEMICAL EQUATIONS

A chemical equation is a theoretical or written representation of what happens during a chemical reaction. The law of conservation of mass states that no atoms can be created or destroyed in a chemical reaction, so the number of atoms that are present in the reactants has to balance the number of atoms that are present in the products. Follow this guide to learn how to balance chemical equations.

Method 1 of 2: Traditional balance

• Write down your given equation



This reaction occurs when propane (C_3H_8) is burned in the presence of oxygen to produce water and carbon dioxide.

• Write down the number of atoms that you have on each side of the equation.

Look at the subscripts next to each atom to find the number of atoms in the equation.

Left side: 3 carbons, 8 hydrogens and 2 oxygens.

Right side: 1 carbon, 2 hydrogens and 3 oxygens.

• Always leave hydrogen and oxygen for last. This means that you will need to balance the carbon atoms first.

• Add a coefficient to the single carbon atom on the right of the equation to balance it with the 3 carbon atoms on the left of the equation.



The coefficient 3 in front of carbon on the right side indicates 3 carbon atoms (just as the subscript 3 on the left side indicates 3 carbon atoms).

In a chemical equation, you can change coefficients, but you should never alter the subscripts.

• Balance the hydrogen atoms next. You have 8 on the left side. So you'll need 8 on the right side.



On the right side, you now added a 4 as the coefficient because the subscript showed that you already had 2 hydrogen atoms.

When you multiply the coefficient 4 times by the subscript 2, you end up with 8.

• The other 6 atoms of Oxygen come from 3 CO_2 subscript 2. (3x2=6 atoms of oxygen+ the other 2= O_2)

• Balance the oxygen atoms.

Because you've added coefficients to the molecules on the right side of the equation, the number of oxygen atoms has changed. You now have 8 oxygen atoms in the water molecule and 6 oxygen atoms in the carbon dioxide molecule. That makes a total of 14 oxygen atoms.

Add a coefficient of 7 to the oxygen molecule on the left side of the equation. You now have 14 oxygen molecules on each side.



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Table of Contents Chemistry Equations Answers Speedy Study Guides Speedy Publishing

1. Understanding the eBook Chemistry Equations Answers Speedy Study Guides Speedy Publishing
 - The Rise of Digital Reading Chemistry Equations Answers Speedy Study Guides Speedy Publishing
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Equations Answers Speedy Study Guides Speedy Publishing
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Equations Answers Speedy Study Guides Speedy Publishing
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Equations Answers Speedy Study Guides Speedy Publishing
 - Personalized Recommendations
 - Chemistry Equations Answers Speedy Study Guides Speedy Publishing User Reviews and Ratings
 - Chemistry Equations Answers Speedy Study Guides Speedy Publishing and Bestseller Lists
5. Accessing Chemistry Equations Answers Speedy Study Guides Speedy Publishing Free and Paid eBooks
 - Chemistry Equations Answers Speedy Study Guides Speedy Publishing Public Domain eBooks
 - Chemistry Equations Answers Speedy Study Guides Speedy Publishing eBook Subscription Services
 - Chemistry Equations Answers Speedy Study Guides Speedy Publishing Budget-Friendly Options
6. Navigating Chemistry Equations Answers Speedy Study Guides Speedy Publishing eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Equations Answers Speedy Study Guides Speedy Publishing Compatibility with Devices
 - Chemistry Equations Answers Speedy Study Guides Speedy Publishing Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Equations Answers Speedy Study Guides Speedy Publishing
 - Highlighting and Note-Taking Chemistry Equations Answers Speedy Study Guides Speedy Publishing
 - Interactive Elements Chemistry Equations Answers Speedy Study Guides Speedy Publishing

8. Staying Engaged with Chemistry Equations Answers Speedy Study Guides Speedy Publishing
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Equations Answers Speedy Study Guides Speedy Publishing
9. Balancing eBooks and Physical Books Chemistry Equations Answers Speedy Study Guides Speedy Publishing
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Equations Answers Speedy Study Guides Speedy Publishing
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Equations Answers Speedy Study Guides Speedy Publishing
 - Setting Reading Goals Chemistry Equations Answers Speedy Study Guides Speedy Publishing
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Equations Answers Speedy Study Guides Speedy Publishing
 - Fact-Checking eBook Content of Chemistry Equations Answers Speedy Study Guides Speedy Publishing
 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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