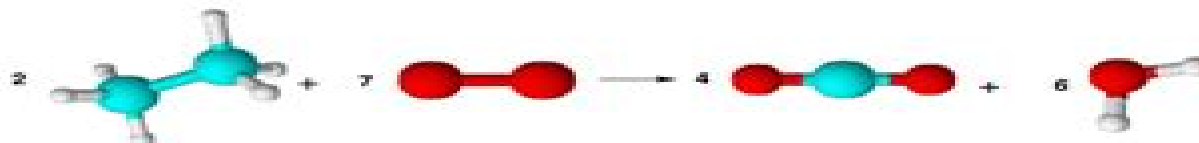


Practice Exercises

- 1.1 (a) SF_6 contains 1 S and 6 F atoms per molecule
 (b) $(\text{C}_2\text{H}_5)_2\text{N}_2\text{H}_2$ contains 4 C, 12 H, and 2 N per molecule
 (c) $\text{Ca}_3(\text{PO}_4)_2$ contains 3 Ca, 2 P, and 8 O atoms per formula unit
 (d) $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ contains 1 Co, 2 N, 12 O, and 12 H per formula unit
- 1.2 (a) NH_4NO_3 contains 2 N nitrogen, 4 H hydrogen, 3 O oxygen atoms per formula unit
 (b) $\text{FeNH}_4(\text{SO}_4)_2$ contains 1 Fe iron, 1 N nitrogen, 4 H hydrogen, 2 S sulfur, 8 O oxygen atoms per formula unit
 (c) $\text{Mo}(\text{NO}_3)_2 \cdot 5\text{H}_2\text{O}$ contains 1 Mo molybdenum, 2 N nitrogen, 11 O oxygen, and 10 H hydrogen atoms per formula unit
 (d) $\text{C}_6\text{H}_6\text{ClNO}_2$ contains 6 C carbon, 4 H hydrogen, 1 Cl chlorine, 1 N nitrogen, and 2 O oxygen atoms per molecule
- 1.3 $\text{C}_2\text{H}_5\text{N}$, CH_3NHCH_3
- 1.4 Reactants: 4 N, 12 H, and 6 O; Products: 4 N, 12 H, and 6 O.
- 1.5 Reactants: 6 N, 42 H, 2 P, 20 O, 3 Ba, 12 C; Products: 3 Ba, 2 P, 20 O, 6 N, 42 H, 12 C; The reaction is balanced.



Review Questions

- 1.1 This answer will be student dependent.
- 1.2 Observation, testing and explanation.
- 1.3 (a) A law is a description of behavior based on the results of many experiments which are true while a theory is a tested explanation of the results of many experiments.
 (b) An observation is a statement that accurately describes something we see, hear, taste, feel or smell while a conclusion is a statement that is based on a series of observations.
 (c) Data are the observations made while performing experiments.
- 1.4 A theory is valid as long as there is no experimental evidence to disprove it. Any experimental evidence that contradicts the theory therefore, disproves the theory.
- 1.5 Matter has mass and occupies space. All items in the question are examples of matter.
- 1.6 A physical change does not change the chemical composition of matter. Melting, boiling, change of shape, or mass, and the formation of a mixture are examples of physical changes to matter.
 A chemical change changes the chemical composition of matter. Formation of new compounds from the reaction of other substances is an example.

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N Noddings



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Table of Contents Chemistry Brady 6th Edition Solutions Manual

1. Understanding the eBook Chemistry Brady 6th Edition Solutions Manual
 - The Rise of Digital Reading Chemistry Brady 6th Edition Solutions Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Brady 6th Edition Solutions Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Brady 6th Edition Solutions Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Brady 6th Edition Solutions Manual
 - Personalized Recommendations
 - Chemistry Brady 6th Edition Solutions Manual User Reviews and Ratings
 - Chemistry Brady 6th Edition Solutions Manual and Bestseller Lists
5. Accessing Chemistry Brady 6th Edition Solutions Manual Free and Paid eBooks
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 - Chemistry Brady 6th Edition Solutions Manual eBook Subscription Services
 - Chemistry Brady 6th Edition Solutions Manual Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Chemistry Brady 6th Edition Solutions Manual Compatibility with Devices
 - Chemistry Brady 6th Edition Solutions Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Chemistry Brady 6th Edition Solutions Manual
 - Interactive Elements Chemistry Brady 6th Edition Solutions Manual
8. Staying Engaged with Chemistry Brady 6th Edition Solutions Manual

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemistry Brady 6th Edition Solutions Manual
- 9. Balancing eBooks and Physical Books Chemistry Brady 6th Edition Solutions Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Brady 6th Edition Solutions Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Brady 6th Edition Solutions Manual
 - Setting Reading Goals Chemistry Brady 6th Edition Solutions Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Brady 6th Edition Solutions Manual
 - Fact-Checking eBook Content of Chemistry Brady 6th Edition Solutions Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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