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Masterton Hurley ISBN 130507937X 9781305079373**

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2 : ATOMS, MOLECULES AND IONS

LECTURE NOTES

Students find this material relatively easy to assimilate; it's almost entirely qualitative. On the other hand, there's a lot of memorizing (sorry, learning) to do. This chapter is coverable in two lectures.

Some general observations:

- ✓ Material in Sections 2.1–2.3 is generally well covered in high-school chemistry courses; no need to dwell on it.
- ✓ Students need to know the molecular formulas of the elements (Figure 2.13), the charges of ions with noble-gas structures and the names and formulas of the common polyatomic ions (Table 2.2). The charges of transition-metal ions will be covered later, in Chapter 4.
- ✓ Naming compounds requires students to distinguish between ionic and molecular substances. It helps to point out that binary molecular compounds are composed of two nonmetals. Almost all ionic compounds contain a metal cation combined with a nonmetal anion or negatively charged polyatomic ion. The flow charts shown in Figures 2.18 and 2.19 should help visual learners.
- ✓ The periodic table will be discussed in greater detail later in the text (Chapter 6).

Lecture 1

I. Atomic Theory

A. Elements

Postulates: Elements consist of tiny particles called atoms, which retain their identity in reactions. In a compound, atoms of two or more elements combine in a fixed ratio of small whole numbers (e.g., 1:1, 2:1, etc.).

B. Components

	relative mass	relative charge	location
proton	1	+1	nucleus
neutron	1	0	nucleus
electron	0.0005	1	outside

Chemistry And Chemical Reactivity 8th Edition Solution Manual

Mark Steyvers



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Table of Contents Chemistry And Chemical Reactivity 8th Edition Solution Manual

1. Understanding the eBook Chemistry And Chemical Reactivity 8th Edition Solution Manual
 - The Rise of Digital Reading Chemistry And Chemical Reactivity 8th Edition Solution Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry And Chemical Reactivity 8th Edition Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry And Chemical Reactivity 8th Edition Solution Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry And Chemical Reactivity 8th Edition Solution Manual
 - Personalized Recommendations
 - Chemistry And Chemical Reactivity 8th Edition Solution Manual User Reviews and Ratings
 - Chemistry And Chemical Reactivity 8th Edition Solution Manual and Bestseller Lists
5. Accessing Chemistry And Chemical Reactivity 8th Edition Solution Manual Free and Paid eBooks
 - Chemistry And Chemical Reactivity 8th Edition Solution Manual Public Domain eBooks
 - Chemistry And Chemical Reactivity 8th Edition Solution Manual eBook Subscription Services
 - Chemistry And Chemical Reactivity 8th Edition Solution Manual Budget-Friendly Options

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6. Navigating Chemistry And Chemical Reactivity 8th Edition Solution Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry And Chemical Reactivity 8th Edition Solution Manual Compatibility with Devices
 - Chemistry And Chemical Reactivity 8th Edition Solution Manual Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry And Chemical Reactivity 8th Edition Solution Manual
 - Highlighting and Note-Taking Chemistry And Chemical Reactivity 8th Edition Solution Manual
 - Interactive Elements Chemistry And Chemical Reactivity 8th Edition Solution Manual
 8. Staying Engaged with Chemistry And Chemical Reactivity 8th Edition Solution Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry And Chemical Reactivity 8th Edition Solution Manual
 9. Balancing eBooks and Physical Books Chemistry And Chemical Reactivity 8th Edition Solution Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry And Chemical Reactivity 8th Edition Solution Manual
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry And Chemical Reactivity 8th Edition Solution Manual
 - Setting Reading Goals Chemistry And Chemical Reactivity 8th Edition Solution Manual
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry And Chemical Reactivity 8th Edition Solution Manual
 - Fact-Checking eBook Content of Chemistry And Chemical Reactivity 8th Edition Solution Manual
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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