

Multiple Choice: (from Zumdahl *Chemistry*, 9th Edition, AP Edition)

Refer to the following diagram for questions 1 – 3

A 100.0 mL Graduated Cylinder is filled with 100.0 mL of water. A few drops of food colouring are added to the top of the water, without mixing. Initially, the top of the graduated cylinder appears to be very dark, while the bottom remains clear and colourless. After 5 min, the liquid in the graduated cylinder appears to be uniformly coloured.



1. Classify the contents in the graduated cylinder before any food colouring is added.

- (A) Element (B) Compound (C) Heterogeneous mixture
(D) Homogeneous mixture

2. Classify the contents in the graduated cylinder just after the food colouring is added.

- (A) Element (B) Compound (C) Heterogeneous Mixture (D) Homogeneous mixture

3. Classify the contents in the graduated cylinder after the 5 minutes has elapsed.

- (A) Element (B) Compound (C) Heterogeneous Mixture (D) Homogeneous mixture

4. A 10.00 g piece of metal is submerged in a graduated cylinder initially containing 20.00 mL of water. With the metal in the cylinder, the level of water is recorded as 21.40 mL. Determine the identity of the metal.

- (A) Al ($\rho = 2.70 \text{ g/mL}$)
(B) Zn ($\rho = 7.13 \text{ g/mL}$)
(C) Ag ($\rho = 10.49 \text{ g/mL}$)
(D) Au ($\rho = 19.32 \text{ g/mL}$)

5. $^{238}_{94}\text{Pu} \rightarrow \text{X} + \alpha$

In the decay represented above, what is the nucleon number of particle X?

- (A) 92 (B) 96 (C) 235 (D) 243

6. Which of the following statements about atoms is CORRECT?

- (A) Atoms are charged because they have different numbers of protons and neutrons
(B) All atoms of an element have the same number of nucleons and electrons.
(C) An atom is mostly filled with protons.
(D) The nucleus is negatively charged.

6. When each of the following is heated to 50°C and the temperature is held constant for 5 min, which one undergoes **only** a physical change.

- (A) egg (B) steak (C) ice cream (D) cake batter

7. A sample of Ti was ionized by electron impact in a mass spectrometer. Information from the mass spectrum about the isotopes of Ti in the sample is shown in Table 2.

Calculate the relative atomic mass of Ti in this sample. Report answers to one decimal place.

Table 1. Results from mass spectrum of Ti sample

m/z	46	47	48	49
Percent Abundance	9.1	7.8	74.6	8.5

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Table of Contents Chemistry 9th Edition Zumdahl Ap Question Answers

1. Understanding the eBook Chemistry 9th Edition Zumdahl Ap Question Answers
 - The Rise of Digital Reading Chemistry 9th Edition Zumdahl Ap Question Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 9th Edition Zumdahl Ap Question Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry 9th Edition Zumdahl Ap Question Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 9th Edition Zumdahl Ap Question Answers
 - Personalized Recommendations
 - Chemistry 9th Edition Zumdahl Ap Question Answers User Reviews and Ratings
 - Chemistry 9th Edition Zumdahl Ap Question Answers and Bestseller Lists

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 - Chemistry 9th Edition Zumdahl Ap Question Answers Public Domain eBooks
 - Chemistry 9th Edition Zumdahl Ap Question Answers eBook Subscription Services
 - Chemistry 9th Edition Zumdahl Ap Question Answers Budget-Friendly Options
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7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry 9th Edition Zumdahl Ap Question Answers
 - Highlighting and Note-Taking Chemistry 9th Edition Zumdahl Ap Question Answers
 - Interactive Elements Chemistry 9th Edition Zumdahl Ap Question Answers
8. Staying Engaged with Chemistry 9th Edition Zumdahl Ap Question Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry 9th Edition Zumdahl Ap Question Answers
9. Balancing eBooks and Physical Books Chemistry 9th Edition Zumdahl Ap Question Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry 9th Edition Zumdahl Ap Question Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry 9th Edition Zumdahl Ap Question Answers
 - Setting Reading Goals Chemistry 9th Edition Zumdahl Ap Question Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry 9th Edition Zumdahl Ap Question Answers
 - Fact-Checking eBook Content of Chemistry 9th Edition Zumdahl Ap Question Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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