

Chemistry 11 Final Examination Review - Answers

Part A - True or False.

Indicate whether each of the following statements is true or false. Correct the false statements.

- F 1. The mass of an electron is equal to the mass of a proton.
The mass of an electron is less than the mass of a proton.
- T 2. The mass of a proton is approximately equal to the mass of a neutron.
- T 3. The atomic number represents the number of protons in a nucleus.
- T 4. The proton has a mass of approximately 1 u.
- F 5. The difference in mass of isotopes of the same element is due to the different number of protons in the nucleus.
The difference in mass of isotopes of the same element is due to the different number of neutrons in the nucleus.
- T 6. The isotope carbon-12 is used as the relative mass standard for the atomic mass scale.
- F 7. The mass of the most-common isotope of each element is listed on the periodic table.
The average mass of the naturally occurring isotopes of each element is listed on the periodic table.

Part B - Multiple Choice

- D 1. A(n) ___ is used to represent a compound.
 a) symbol b) equation c) subscript d) formula
- A 2. ___ atoms or groups of atoms are called ions.
 a) charged b) diatomic c) neutral d) monatomic
- A 3. For the formula of a compound to be correct, the algebraic addition of the charges on the atoms or ions in the compound must add up to ____.
 a) zero b) one c) two d) four
- D 4. Potassium bromide is an example of a(n) ___ compound.
 a) molecular b) organic c) polyatomic d) ionic
- B 5. The only common polyatomic ion that has a positive charge is the ___ ion.
 a) phosphate b) ammonium c) sulfate d) nitrate
- A 6. In the formula H_2SO_4 , the number 4 would be called a(n) ____.
 a) subscript b) oxidation number c) coefficient d) charge
- B 7. Which subatomic particle contributes the least to the mass of an atom?
 a) nucleon b) electron c) proton d) neutron
- D 8. In the free, or uncombined, state the number of protons in the nucleus of an element must equal the ____.
 a) mass number b) number of neutrons in the nucleus c) mass number - atomic number d) number of electrons present

Chemistry 11 Final Exam Review

John T. Moore



Chemistry 11 Final Exam Review:

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