

Chemische binding

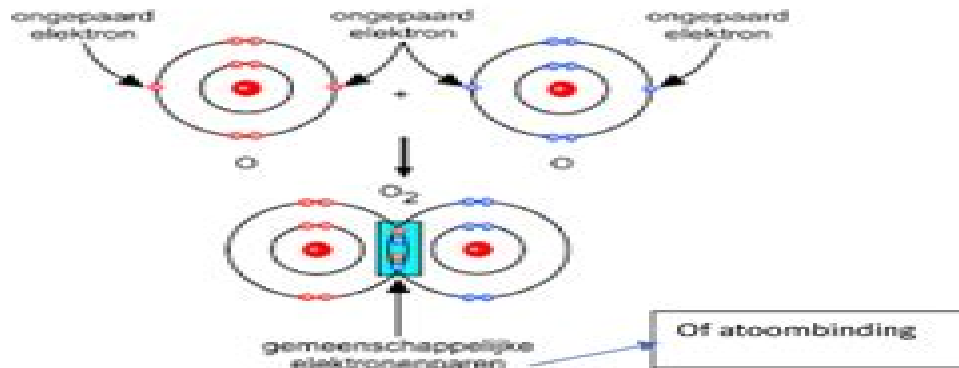
Wat is een covalente binding of atoombinding?

Atoombinding is een binding van twee of meer niet-metalen. Wanneer atomen een of meer elektronen gemeenschappelijk stellen om de edelgasconfiguratie te bereiken, ontstaat er een covalente binding of atoombinding. Een atoombinding is zwakker dan een ion binding.

Enkelvoudige atoombinding

Niet-metalen reageren vaak met andere niet-metalen.

Niet-metalen hebben in hun buitenste schil 5, 6 of 7 elektronen: ze vertonen dus de neiging om 1 of meer elektronen op te nemen. Bij een reactie tussen twee niet-metalen zullen er dus geen elektronen overgedragen worden. Er zullen dus geen ionen ontstaan. Beide niet-metalen vertonen immers de neiging om elektronen op te nemen, maar geen van beide vertoont de neiging om elektronen af te geven. Toch kunnen de niet-metalen een stabiele edelgasconfiguratie verkrijgen.



Dit kan je ook noteren met de lewisnotatie (kijk boek p. 96)

Meervoudige covalente bindingen

Soms worden twee of drie elektronenparen gemeenschappelijk gesteld, als atomen twee of drie valentie-elektronen te kort hebben voor een edelgasstructuur. Je hebt een dubbele covalente binding en een drievoudige covalente binding. (kijk blz 97 voor lewisnotatie)

Covalente binding tussen atomen van verschillende soorten

Covalente binding tussen atomen van verschillende soorten gebeuren op dezelfde manier als bij atomen van dezelfde soort, maar hier speelt een bijkomende fenomeen een rol: de elektronegativiteit of **elektronegative waarde (ENW)**.

Chemische Binding Als Electrostatisch Verschijnsel

**Anton Eduard van Arkel, Jan Hendrik
de Boer**



Chemische Binding Als Electrostatisch Verschijnsel:

An Introduction to the Chemistry of Complex Compounds Aleksander Abramovich Grinberg, 2013-10-22 An Introduction to the Chemistry of Complex Compounds discusses the fundamental concepts that are essential in understanding the underlying principles of complex compounds The coverage of the book includes the compounds of the hexa penta and tetrammine type compounds of the tri dl monoamine and hexacido types for the coordination number of 6 and complex compounds with a coordination number of 4 The text also covers the effects and chemical properties of complex compounds such as the nature of the force of complex formation the mutual effects of coordinated groups and acid base properties oxidation reduction properties and solution equilibriums of complex compounds The book will be of great use to chemists and chemical engineers **Bulletin** Electrochemical Society, 1929 Chemisch weekblad , 1927 Includes section

Boekaankondigingen **Physical and Chemical Aspects of Adsorbents and Catalysts** Jan Hendrik Boer, 1970

Nederlandsch tijdschrift voor geneeskunde , 1929 **Crystal Structure and Chemical Bonding in Inorganic Chemistry** Albrecht Rabenau, Coenraad Jozef Maria Rooymans, 1975 **Natuurwetenschappelijk tijdschrift** , 1926

Chemische binding als electrostatisch verschijnsel Anton Eduard van Arkel, Jan Hendrik de Boer, 1930 **Onderzoek in opdracht** L.J. Dorsman, P.J. Knegtmans, 2007 **New Dictionary of Scientific Biography** Noretta Koertge, 2008 Also available online as part of the Gale Virtual Reference Library under the title Complete dictionary of scientific biography

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Table of Contents Chemische Binding Als Electrostatistisch Verschijnsel

1. Understanding the eBook Chemische Binding Als Electrostatistisch Verschijnsel
 - The Rise of Digital Reading Chemische Binding Als Electrostatistisch Verschijnsel
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemische Binding Als Electrostatistisch Verschijnsel
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemische Binding Als Electrostatistisch Verschijnsel
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemische Binding Als Electrostatistisch Verschijnsel
 - Personalized Recommendations
 - Chemische Binding Als Electrostatistisch Verschijnsel User Reviews and Ratings
 - Chemische Binding Als Electrostatistisch Verschijnsel and Bestseller Lists

5. Accessing Chemische Binding Als Electrostatisch Verschijnsel Free and Paid eBooks
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 - Chemische Binding Als Electrostatisch Verschijnsel eBook Subscription Services
 - Chemische Binding Als Electrostatisch Verschijnsel Budget-Friendly Options
6. Navigating Chemische Binding Als Electrostatisch Verschijnsel eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemische Binding Als Electrostatisch Verschijnsel Compatibility with Devices
 - Chemische Binding Als Electrostatisch Verschijnsel Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemische Binding Als Electrostatisch Verschijnsel
 - Highlighting and Note-Taking Chemische Binding Als Electrostatisch Verschijnsel
 - Interactive Elements Chemische Binding Als Electrostatisch Verschijnsel
8. Staying Engaged with Chemische Binding Als Electrostatisch Verschijnsel
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemische Binding Als Electrostatisch Verschijnsel
9. Balancing eBooks and Physical Books Chemische Binding Als Electrostatisch Verschijnsel
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemische Binding Als Electrostatisch Verschijnsel
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemische Binding Als Electrostatisch Verschijnsel
 - Setting Reading Goals Chemische Binding Als Electrostatisch Verschijnsel
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemische Binding Als Electrostatisch Verschijnsel
 - Fact-Checking eBook Content of Chemische Binding Als Electrostatisch Verschijnsel
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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