

molecular equation:



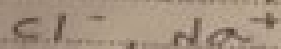
complete ionic equation:



net ionic equation:



spectator ions:



general reaction type:

Double replacement/displacement

#### VI. Heating Cu with Atmospheric O<sub>2</sub>

appearance the unreacted Cu mesh:

shiny copper

appearance of the Cu mesh after heating and cooling:

dull dark brown

chemical equation:



general reaction type:

combination

Which species are oxidized?

Cu

Which species are reduced?

O

#### VII. Reacting CuSO<sub>4</sub> Solution with Steel Wool (Fe)

appearance of the steel wool (Fe) before reaction:

silver solid

appearance of the CuSO<sub>4</sub> solution before reaction:

light blue

appearance of the steel wool (Fe) after the reaction has occurred:

dark brown solid

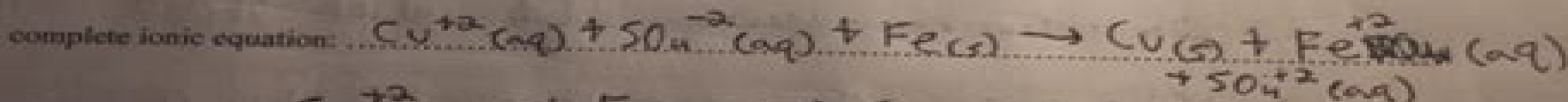
appearance of the CuSO<sub>4</sub> solution after the reaction has occurred:

light blue liquid

molecular equation:



complete ionic equation:



net ionic equation:



spectator ions:



general reaction type:

single displacement

Which species are oxidized?

Fe

Which species are reduced?

Cu<sup>+2</sup>

# Chemical Reactions Andequations Lab Manual Answers

**Neena Sinha, R.Rangarajan, Rajesh  
Kumar**



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*Laboratory Manual for Principles of General Chemistry* J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section

*Lab Manual for General, Organic, and Biochemistry* Denise Guinn, Rebecca Brewer, 2009-08-21 Teaching all of the necessary concepts within the constraints of a one term chemistry course can be challenging Authors Denise Guinn and Rebecca Brewer have drawn on their 14 years of experience with the one term course to write a textbook that incorporates biochemistry and organic chemistry throughout each chapter emphasizes cases related to allied health and provides students with the practical quantitative skills they will need in their professional lives Essentials of General Organic and Biochemistry captures student interest from day one with a focus on attention getting applications relevant to health care professionals and as much pertinent chemistry as is reasonably possible in a one term course Students value their experience with chemistry getting a true sense of just how relevant it is to their chosen profession To browse a sample chapter view sample ChemCasts and more visit [www.whfreeman.com/gob](http://www.whfreeman.com/gob)

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