Worksheet: Chemical Equations Review

Name: Date: Period:

Sections 1: Identify the type of reaction

For the following reactions, indicate whether the following are examples of synthesis, decomposition, combustion, single displacement, or double displacement.

- Na₃PO₄ + 3KOH → 3 NaOH + K₃PO₄
- MgCl₂ +Li₂CO₃ → MgCO₃ + 2 LiCl
- C₆H₁₂ + 9O₂ → 6CO₂ +6 H₂O ______
- 4. Pb + FeSO₄ → PbSO₄ + Fe
- 5. CaCO₃ → CaO + CO₂
- 6. P₄ + 3O₂ → 2P₂O₃
- 8. 2AaNO₃ + Cu → Cu(NO₃)₂ + 2Ag
- 9. C₃H₆O + 4O₂ → 3CO₂ + 3 H₂O
- 10.2C₅H₅ + Fe → Fe(C₅H₅)₂
- 11. SeCl₆ + O₂ → SeO₂ + 3Cl₂
- 12.2 Mgl₂ + Mn(SO₃)₂ → 2MgSO₃ + MnI₄
- 13.O₃ → O + O₂ _____
- 14.2NO₂ → 2O₂ + N₂_____

Section 2: Practicing Equation Balancing

Balance each of the following equations.

- _C₆H₆ + _O₂ → _H₂O + _CO₂
- _Nal + _Pb(SO₄)₂ → _Pbl₄ + _Na₂SO₄
- NH₃ + _O₂ → _NO + _H₂O
- Fe(OH)₃ → Fe₂O₃ + H₂O
- 5. $_HNO_3 + _Mg(OH)_2 \rightarrow _H_2O + _Mg(NO_3)_2$

Chemical Equations And Reactions Section 1 Answers

T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder

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