

Acid-Base Titrations Pre-Lab Questions (10 Points)

Lab Instructor	Date		
. In an experiment, 39.26	ce provided or on a separate s	on was required to ti	trans 37.54 mL of
density to be 1.00 p/ml	fusion to a phenolphthalein end If the percent (by weight) of ace beog. (Ma) (We) = (Ma) (We	tic acid in the solution	molarity of the on (assuming its
	H2 (57-54m1) = (0.10)	(39.26)	M1)
	H2 (57-54m1) = (0.10) HL = 4.165	24 PM (39.24 14 - 0.1110	m1)

- 2. What is the proper disposal of all chemicals in this experiment?

 Discord of an solutions, Her solution, rinsing solutions, and thiration solutions into your maste beauting. It your maste beautingers fixed, displayed of all solutions in the waste out base comb ben container.
- 3. List some the hazards associated with sodium hydroxide.

 Sodium hydroxide is to xic and, composite, and can easily absorbed by the sun.

Chemfax Acid Base Titrations Lab Prelab Answers

Sten Johansson

Chemfax Acid Base Titrations Lab Prelab Answers:

<u>Identification of Mis</u>conceptions Acid-base Titrations in Nonaqueous Solvents James Sherwood Fritz, 1952 Concerning Acid-base Titrations Lockwood Ryerson Gibbs, 1995 Agueous Acid-base Equilibria and Titrations Robert De Levie, 1999 This book will give students a thorough grounding in pH and associated equilibria material absolutely fundamental to the understanding of many aspects of chemistry It is in addition a fresh and modern approach to a topic all too often taught in an out moded way This book uses new theoretical developments which have led to more generalized approaches to equilibrium problems these approaches are often simpler than the approximations which they replace Acid base problems are readily addressed in terms of the proton condition a convenient amalgam of the mass and charge constraints of the chemical system considered The graphical approach of Bjerrum Hagg and Sillen is used to illustrate the orders of magnitude of the concentrations of the various species involved in chemical equilibria Based on these concentrations the proton condition can usually be simplified often leading directly to the value of the pH In the description of acid base titrations a general master equation is developed It provides a continuous and complete description of the entire titration curve which can then be used for computer based comparison with experimental data Graphical estimates of the steepness of titration curves are also developed from which the practicality of a given titration can be anticipated Activity effects are described in detail including their effect on titration curves. The discussion emphasizes the distinction between equilibrium constants and electrometric pH measurements which are subject to activity corrections and balance equations and spectroscopic pH measurements which are not Finally an entire chapter is devoted to what the pH meter measures and to the experimental and theoretical uncertainties involved Acid/Base Titrations Michael O'Brien, 1999 **Numerical Evaluation of Potentiometric Acid-base Titrations** Sten Johansson, 1978 PH, Buffers and Acid-base Titrations LaMotte Chemical Products Co. Educational Products Division.197?

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